

Introduction 11

An Electron's Address

Name: _____
 Date: _____
 Class: _____

Introduction: Energy Levels and Sublevels

In addition to the energy level model that you used last semester, scientists have developed another model. This system models orbitals, which are regions in space where electrons

reside. Electrons spend most of their energy levels and sublevels. Scientists have found that electrons spend most of their energy levels and sublevels. Design the following table:

Table 1

Energy Level	Number of sublevels in this energy level
1	1
2	2
3	3
4	4

How does this model describe energy as a structural model of the 1^{st} energy level because there are only one sublevel in this energy level?

Critical Thinking Questions

- How many sublevels exist in the 1^{st} energy level?
 The sublevel in the 1^{st} energy level is the only sublevel in this energy level.
- How many sublevels exist in the 2^{nd} energy level?
 The sublevels in the 2^{nd} energy level are the s and p sublevels.
- How many sublevels exist in the 3^{rd} energy level?
 The sublevels in the 3^{rd} energy level are the s , p , and d sublevels.
- How many sublevels would you expect to have for the 4^{th} energy level?
 The number of sublevels I would expect to have for the 4^{th} energy level would be 4 sublevels. There would be 4 sublevels in the 4^{th} energy level.
- How many sublevels would you expect to have for the 5^{th} energy level?
 The number of sublevels I would expect to have for the 5^{th} energy level would be 5 sublevels. There would be 5 sublevels in the 5^{th} energy level.

Chemquest 11 Answers

RJ Alexander



Chemquest 11 Answers:

The book delves into Chemquest 11 Answers. Chemquest 11 Answers is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Chemquest 11 Answers, encompassing both the fundamentals and more intricate discussions.

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