

# OPERATIONS MANAGEMENT

JOHN NAYLOR



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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  4 Al + 3  $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g Al  $\times \frac{1 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}} \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$  4 mol Al / 1 mol  $\text{Al}_2\text{O}_3$  c determine the percent yield yield actual  $\times 100$  260.0 g  $\times 100$  86.01 theoretical 302.3 g

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