

### Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility:** the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.  
**saturated solution:** a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution:** a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible:** describes liquids that dissolve in each other. **immiscible:** describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No, if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47.  $5.55 \times 10^2$  g  $\text{AgNO}_3$
48. Solubility increases with pressure.
49. a.  $1.6 \times 10^{-2}$  g/L  
b.  $4.7 \times 10^{-2}$  g/L
50. Dilute and concentrated are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.  
a. 1.5M KCl  
b.  $5.3 \times 10^{-2}$  M  $\text{MgCl}_2$
52.  $2.00 \times 10^2$  mol
53. a.  $5.0 \times 10^{-2}$  mol NaCl, 20 g NaCl  
b. 1.0 mol  $\text{KNO}_3$ ,  $1.0 \times 10^2$  g  $\text{KNO}_3$   
c.  $2.5 \times 10^2$  mol  $\text{CaCl}_2$ , 2.0 g  $\text{CaCl}_2$
54. a.  $2.3 \times 10^2$  g NaCl  
b. 2.0 g  $\text{MgCl}_2$
55. a. 10% (v/v) ethanol  
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water  
b. 1.5M  $\text{KNO}_3$   
c. 0.100M  $\text{MgCl}_2$
58. The effective molality of the  $\text{Ca}(\text{NO}_3)_2$  solution is 3m. The effective molality of the NaCl solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g  $\text{H}_2\text{O}$  to 32.0 g  $\text{C}_2\text{H}_5\text{OH}$ .
63. a. 100.26°C  
b. 101.54°C
64. a.  $-4.40^\circ\text{C}$   
b.  $-2.2^\circ\text{C}$
65. a.  $-1.1^\circ\text{C}$   
b.  $-0.24^\circ\text{C}$   
c.  $-1.3^\circ\text{C}$

### Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

# Prentice Hall Chemistry Chapter 13 Assessment Answers

**Burns**



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*Chemical Property Estimation* Edward Baum, 2018-05-11 Our world is widely contaminated with damaging chemicals and companies create thousands of new potentially dangerous chemicals each year Due to the difficulty and expense of obtaining accurate measurements and the unreliability of reported values we know surprisingly little about the properties of these contaminants Determining the properties of chemicals is critical to judging their impact on environmental quality and in making decisions about emission rates clean up and other important public health issues *Chemical Property Estimation* describes modern methods of estimating chemical properties methods which cost much less than traditional laboratory techniques and are sufficiently accurate for most environmental applications Estimation methods are used to screen chemicals for testing design monitoring and analysis methods design clean up procedures and verify experimental measurements The book discusses key methods for estimating chemical properties and considers their relative strengths and weaknesses Several chapters are devoted to the partitioning of chemicals between air water soil and biota and properties such as solubility vapor pressure and chemical transport Each chapter begins with a review of relevant theory and background information explaining the applications and limitations of each method Sample calculations and practical advice on how and when to use each method are included as well Each method is evaluated for accuracy and reliability Computer software databases and internet resources are evaluated as well as other supplementary material such as fundamental constants units of measure and more    **Introductory Chemistry** Nivaldo J. Tro, 2006    Sm Fundamentals Chemistry T/I Burns, 1999-03    *High Temperature Corrosion* César A. C. Sequeira, 2018-12-14 Reviews the science and engineering of high temperature corrosion and provides guidelines for selecting the best materials for an array of system processes High temperature corrosion HTC is a widespread problem in an array of industries including power generation aerospace automotive and mineral and chemical processing to name a few This book provides engineers physicists and chemists with a

balanced presentation of all relevant basic science and engineering aspects of high temperature corrosion It covers most HTC types including oxidation sulfidation nitridation molten salts fuel ash corrosion H<sub>2</sub>S H<sub>2</sub> corrosion molten fluoride HF corrosion and carburization It also provides corrosion data essential for making the appropriate choices of candidate materials for high temperature service in process conditions A form of corrosion that does not require the presence of liquids high temperature corrosion occurs due to the interaction at high temperatures of gases liquids or solids with materials HTC is a subject of increasing importance in many areas of science and engineering and students researchers and engineers need to be aware of the nature of the processes that occur in high temperature materials and equipment in common use today especially in the chemical gas petroleum electric power metal manufacturing automotive and nuclear industries Provides engineers and scientists with the essential data needed to make the most informed decisions on materials selection Includes up to date information accompanied by more than 1 000 references 80% of which from within the past fifteen years Includes details on systems of critical engineering importance especially the corrosion induced by low energy radionuclides Includes practical guidelines for testing and research in HTC along with both the European and International Standards for high temperature corrosion engineering Offering balanced in depth coverage of the fundamental science behind and engineering of HTC High Temperature Corrosion Fundamentals and Engineering is a valuable resource for academic researchers students and professionals in the material sciences solid state physics solid state chemistry electrochemistry metallurgy and mechanical chemical and structural engineers *Resources in education* ,1984-09 Prentice Hall Health's Q and A Review of Medical Technology/clinical Laboratory Science Anna P. Ciulla,Georganne K. Buescher,2002 A valuable review for a wide range of laboratory professionals this book prepares candidates for certification examinations by presenting them with the latest technology and terminology as well as current test taking formats Its large number of practice questions variety of practice modes and explanations for clarification prepare learner for success on examinations Comprehensive coverage of laboratory medicine includes clinical chemistry hematology hemostasis immunology immunochemistry microbiology urinalysis and body fluids molecular diagnostics laboratory calculations general laboratory principles and safety laboratory management education and computers and laboratory informatics **Water Resources Engineering** Larry W. Mays,2019-04-02 Modern water conveyance and storage techniques are the product of thousands of years of human innovation today we rely on that same innovation to devise solutions to problems surrounding the rational use and conservation of water resources with the same overarching goal to supply humankind with adequate clean freshwater Water Resources Engineering presents an in depth introduction to hydrological and hydraulic processes with rigorous coverage of both core principles and practical applications The discussion focuses on the engineering aspects of water supply and water excess management relating water use and the hydrological cycle to fundamental concepts of fluid mechanics energy and other physical concepts while emphasizing the use of up to date analytical tools and methods Now in

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