

Chapter 4 – Atomic Structure

Short Answer: Answer the following questions in complete sentences using science vocabulary. Make sure that your answers are complete and detailed.

1. A sample of calcium contains calcium-40, calcium-44, calcium-42, calcium-48, calcium-43, and calcium-46 atoms. Explain why these atoms have different mass numbers but must have the same atomic number.

The atoms have different mass numbers because they have a different number of neutrons. The mass number is equaled to the number of protons plus neutrons.

The atoms must have the same atomic number because the atomic number is equaled to the number of protons. The number of protons is unique for each element.

2. Complete the following table. Circle the name of the element(s) that are isotopes. Complete Bohr's Model diagrams for the element with eleven electrons and the element with an atomic number of 9.

Element Name	Element Symbol	Atomic Number	Mass Number	Protons	Neutrons	Electrons
Scandium	Sc	21	45	21	24	21
Fluorine	F	9	19	9	10	9
Sodium	Na	11	24	11	13	11
Answers Vary					82	

Chapter 4 Atomic Structure Answers

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The Science For Conservators Series The Conservation Unit Museums and Galleries Commission, 2008-05-20 For more than ten years The Science for Conservators Series has provided the key basic texts for conservators throughout the world Scientific concepts are basic to the conservation of artefacts of every type yet many conservators have little or no scientific training These introductory volumes provide non scientists with the essential theoretical background to their work

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Foundations of College Chemistry, Student Solutions Manual Morris Hein, Susan Arena, Cary Willard, 2022-03-09

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applications Finally the topics related to safety health and societal impact of nanotechnology are discussed

McGraw-Hill's SAT Subject Test Chemistry, 3rd Edition Thomas Evangelist,2011-02-03 Expert guidance on the Chemistry exam Many colleges and universities require you to take one or more SAT II Subject Tests to demonstrate your mastery of specific high school subjects McGraw Hill s SAT Subject Test Chemistry is written by experts in the field and gives you the guidance you need perform at your best This book includes 4 full length sample tests updated for the latest test format 40 top tips to remember on test day Glossary of tested chemistry terms and formulas Tips and strategies from one of the most popular teachers at the renowned Brooklyn Technical High School Diagnostic test to pinpoint strengths and weaknesses Step by step review of all topics covered on the exam In depth coverage of the lab experiment questions that are a major test feature Charts tables and illustrations to simplify and reinforce learning Practice tests just like the real SAT Subject Test in Chemistry Test taking tips and strategies

Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction

The Mining Magazine ,1927

On the Application of the Quantum Theory to Atomic Structure Niels Bohr,1924

Study Guide with Answers to Selected Problems H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

MATERIALS SCIENCE AND ENGINEERING, SIXTH EDITION V. RAGHAVAN,2015-05-01 This well established and widely adopted book now in its Sixth Edition provides a thorough analysis of the subject in an easy to read style It analyzes systematically and logically the basic concepts and their applications to enable the students to comprehend the subject with ease The book begins with a clear exposition of the background topics in chemical equilibrium kinetics atomic structure and chemical bonding Then follows a detailed discussion on the structure of solids crystal imperfections phase diagrams solid state diffusion and phase transformations This provides a deep insight into the structural control necessary for optimizing the various properties of materials The mechanical properties covered include elastic anelastic and viscoelastic behaviour plastic deformation creep and fracture phenomena The next four chapters are devoted to a detailed description of electrical conduction superconductivity semiconductors and magnetic and dielectric properties The final chapter on Nanomaterials is an important addition to the sixth edition It describes the state of art developments in this new field This eminently readable and student friendly text not only provides a masterly analysis of all the relevant topics but also makes them comprehensible to the students through the skillful use of well drawn diagrams illustrative tables worked out examples and in many other ways The book is primarily intended for

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