

CHEMISTRY

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PERIODIC TABLE OF THE ELEMENTS

The periodic table is color-coded by groups:

- 1 (pink):** s-subshell, 1 orbital
- 2 (orange):** s-subshell, 1 orbital
- 3-10 (purple):** d-subshell, 5 orbitals
- 11-18 (green):** p-subshell, 3 orbitals
- 19-20 (blue):** s-subshell, 1 orbital
- 21-32 (yellow):** d-subshell, 5 orbitals
- 33-36 (light blue):** p-subshell, 3 orbitals
- 37-54 (light purple):** s-subshell, 1 orbital
- 55-86 (light green):** p-subshell, 3 orbitals
- 87-118 (light blue):** s-subshell, 1 orbital



ATOMIC STRUCTURE

Atomic Number, Z , is of protons in the nucleus. For neutral atom, $Z = +$ of electrons. For ions of charge, $+Q$, $Z = +Q$ (cation) or $Z = -Q$ (anion). Atomic mass, A , is the sum of protons and neutrons in the nucleus. Isotopes atoms with the same Z , different A . The atomic mass has a number of protons and neutrons. A number of isotopes is practical with use the atomic mass. Number of Atomic Weight is practical average of atomic masses for a given element.

NUCLEAR CHEMISTRY

Nuclear reactions with the nucleus particles consist of alpha, beta, gamma, neutron and positron. Energy (E) is related to mass (m) by the Einstein equation, $E = mc^2$. Nuclear reaction that changes Z , number is symbolized by α , β , γ , n , p or e^+ (eg. α -decay, β -decay, γ -ray, n -ray, p -ray, e^+ -ray).

Types of processes:
 Transmutation: Two elements made from particle collisions.
 Fusion: Small nuclei combine to form a larger nucleus.
 Fission: Large nuclei split into smaller nuclei.
 Radioactive decay:
 α -decay: α particle is emitted.
 β -decay: β particle is emitted.
 γ -decay: γ ray is emitted.
 Positron emission: e^+ is emitted.
 Electron capture: e^- is captured.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Quantum numbers describe the energy, size, and shape of an orbital. The principal quantum number (n) is the most important. It determines the energy level and the size of the orbital. The azimuthal quantum number (l) determines the shape of the orbital. The magnetic quantum number (m_l) determines the orientation of the orbital. The spin quantum number (m_s) determines the spin of the electron in the orbital.

Orbitals are labeled by their principal quantum number (n) and azimuthal quantum number (l). For example, a 2p orbital has $n=2$ and $l=1$. The number of orbitals in a shell is n^2 . The number of orbitals in a subshell is $2l+1$.

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the behavior of electrons in atoms. The Pauli Exclusion Principle states that no two electrons can have the same set of quantum numbers. The Aufbau Principle states that electrons fill orbitals in order of increasing energy. The Hund's Rule states that electrons fill degenerate orbitals singly before pairing up.

n	l	m _l	m _s	Orbital
1	0	0	±1/2	1s
2	0	0	±1/2	2s
2	1	0	±1/2	2p _z
2	1	±1	±1/2	2p _x , 2p _y
3	0	0	±1/2	3s
3	1	0	±1/2	3p _z
3	1	±1	±1/2	3p _x , 3p _y
3	2	0	±1/2	3d _{z²}
3	2	±1	±1/2	3d _{xy} , 3d _{xz} , 3d _{yz}

Atomic configuration: Shows the arrangement of electrons in an atom. The Pauli Exclusion Principle states that no two electrons can have the same set of quantum numbers. The Aufbau Principle states that electrons fill orbitals in order of increasing energy. The Hund's Rule states that electrons fill degenerate orbitals singly before pairing up.

The energy level diagram shows the relative energies of atomic orbitals for different principal quantum numbers (n). The energy increases with n . For a given n , the energy order is $s < p < d < f$. The diagram illustrates the filling order of orbitals according to the Aufbau principle.

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