

## Solutions to Chen's Plasma Physics

Kalman Krizhník

**1-1.** Compute the density (in units of  $m^{-3}$ ) of an ideal gas under the following conditions:

a) At  $0^\circ\text{C}$  and 760 Torr pressure (1 Torr = 1mm Hg). This is called the Loschmidt number.

b) In a vacuum at  $10^{-3}$  Torr at room temperature ( $20^\circ\text{C}$ ). This number is a useful one for the experimentalist to know by heart ( $10^{-3}$  Torr = 1 micron).

a) Avogadro's number is  $N_A = 6.022 \times 10^{23}$ . One mole of gas at STP occupies 22.4 liters. 1 liter is  $1 \times 10^{-3}$  cubic meters. Thus, the number per cubic meter is  $N_A/n = 6.022 \times 10^{23}/(22.4 \times 10^{-3}) = 2.66 \times 10^{25} m^{-3}$ . Thus, the Loschmidt number is  $2.66 \times 10^{25} \square$

b) Using  $PV=NkT$ , we obtain (with  $R=1.4 \times 10^{-23} J K^{-1}$  and 1 Torr = 133 Pa):

$$n = \frac{N}{V} = \frac{P}{kT} = \frac{10^{-3} \times 133}{1.4 \times 10^{-23} \times (20 + 273)} = 3.3 \times 10^{19} m^{-3} \quad \square \quad (1)$$

**1-2.** Derive the constant  $A$  for a normalized one-dimensional Maxwellian distribution

$$f(u) = Ae^{-mu^2/2kT} \quad (2)$$

such that

$$\int_{-\infty}^{\infty} f(u)du = 1 \quad (3)$$

This one is straightforward. Just integrate:

$$1 = \int_{-\infty}^{\infty} Ae^{-mu^2/2kT} du = A\sqrt{\frac{2\pi kT}{m}} \Rightarrow A = \sqrt{\frac{m}{2\pi kT}} \quad \square \quad (4)$$

**1-4.** Compute the pressure, in atmospheres and in  $tons/ft^2$ , exerted by a thermonuclear plasma on its container. Assume  $kT_e = kT_i = 20keV$ ,  $n = 10^{21}m^{-3}$  and  $p = nkT$ , where  $T = T_i + T_e$ .

This is just unit conversion, albeit with units that nobody really ever remembers. For reference,  $1 keV = 1.6 \times 10^{-19}J$ , so

$$p = 10^{21} \times (20keV + 20keV) = 4 \times 10^{22}m^{-3} keV = 4 \times 10^3m^{-3} J = 4 \times 10^5 N/m^2 \quad (5)$$

But  $1 atm = 10^5 N/m^2 = 1 ton/ft^2$ , so (Note: I think there is a mistake in Chen's solutions here. If I am mistaken, please let me know).

$$p = 0.04 atm = 0.04 ton/ft^2 \quad \square \quad (6)$$

**1-5.** In a strictly steady state situation, both the ions and the electrons will follow the Boltzmann relation

$$n_j = n_0 e^{-q_j \phi / kT_j} \quad (7)$$

For the case of an infinite, transparent grid charged to a potential  $\phi$ , show that the shielding distance is given approximately by

$$\lambda_D^{-2} = \frac{ne^2}{\epsilon_0} \left( \frac{1}{kT_e} + \frac{1}{kT_i} \right) \quad (8)$$

Show that  $\lambda_D$  is determined by the temperature of the colder species.

We'll use Poisson's equation

$$\nabla^2 \phi = \frac{en_e - en_i}{\epsilon_0} = \frac{en_0}{\epsilon_0} (e^{e\phi/kT_e} - e^{-e\phi/kT_i}) \approx \frac{en_0}{\epsilon_0} \left( 1 + \frac{e\phi}{kT_e} - 1 - \frac{-e\phi}{kT_i} \right) = \frac{e^2 n}{\epsilon_0} \left( \frac{\phi}{kT_e} + \frac{\phi}{kT_i} \right) \quad (9)$$

# Chen Plasma Physics Solutions

**M Mosston**



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web apr 7 2014 molarity moles of solute liters of solution  $8.4 \text{ L} \times 2.2 \text{ M} = 18.48 \text{ mol}$  a first convert 250 ml to liters  $250 \text{ mL} = 0.25 \text{ L}$  then calculate molarity  $5 \text{ moles} / 0.25 \text{ liters} = 20 \text{ M}$  c a solution with molarity 2 requires 2 M of NaOH per liter so  $4 \times 2.8 \text{ M} = 11.2 \text{ M}$  a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture can be calculated using the following formula  $M_{\text{mix}} = \frac{m_1 + m_2}{V_1 + V_2}$   $\times 1000$   $\times 1.5 \times 480 + 1.2 \times 520 = 480 + 624 = 1104$   $\times 1000 = 1.104 \text{ M}$  q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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web 1 calculate the molarity of a solution which contains 0.40 mol of a substance dissolved in 1.6 L of a solution 2 what is the molarity of a solution containing 325 g of NaCl dissolved in 750 ml of solution 3 140 g of KCl is dissolved in 600 ml of water what is the molarity 4 724 g of ammonium phosphate in 4500 ml of alcohol

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web determine the molarity for each of the following solutions 0.444 mol of  $\text{COCl}_2$  in 0.654 L of solution 98.0 g of phosphoric acid  $\text{H}_3\text{PO}_4$  in 1.00 L of solution 0.2074 g of calcium hydroxide  $\text{Ca(OH)}_2$  in 40.00 ml of solution 10.5 kg

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web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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