

## Section 12.1

### OBJECTIVE 1

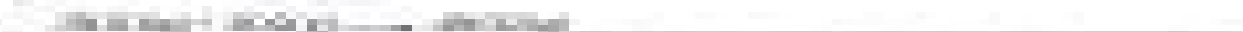
Write complete balanced redox reactions that are indicated.

1.  $\text{Fe}^{2+} + \text{MnO}_4^- + \text{H}^+$ 
  - (a) Write the balanced redox reaction that occurs in acidic solution. In your answer, include the balanced equation, the number of electrons transferred, and the number of  $\text{H}^+$  ions that are consumed or produced.
  - (b) Write the balanced redox reaction that occurs in basic solution. In your answer, include the balanced equation, the number of electrons transferred, and the number of  $\text{OH}^-$  ions that are consumed or produced.
2.  $\text{Fe}^{2+} + \text{MnO}_4^- + \text{H}^+$ 
  - (a) Write the balanced redox reaction that occurs in acidic solution. In your answer, include the balanced equation, the number of electrons transferred, and the number of  $\text{H}^+$  ions that are consumed or produced.
  - (b) Write the balanced redox reaction that occurs in basic solution. In your answer, include the balanced equation, the number of electrons transferred, and the number of  $\text{OH}^-$  ions that are consumed or produced.
3.  $\text{Fe}^{2+} + \text{MnO}_4^- + \text{H}^+$ 
  - (a) Write the balanced redox reaction that occurs in acidic solution. In your answer, include the balanced equation, the number of electrons transferred, and the number of  $\text{H}^+$  ions that are consumed or produced.
  - (b) Write the balanced redox reaction that occurs in basic solution. In your answer, include the balanced equation, the number of electrons transferred, and the number of  $\text{OH}^-$  ions that are consumed or produced.
4.  $\text{Fe}^{2+} + \text{MnO}_4^- + \text{H}^+$ 
  - (a) Write the balanced redox reaction that occurs in acidic solution. In your answer, include the balanced equation, the number of electrons transferred, and the number of  $\text{H}^+$  ions that are consumed or produced.
  - (b) Write the balanced redox reaction that occurs in basic solution. In your answer, include the balanced equation, the number of electrons transferred, and the number of  $\text{OH}^-$  ions that are consumed or produced.

5. Write the balanced redox reaction that occurs in acidic solution.



a. Balance the equation:



b. Write the balanced redox reaction that occurs in acidic solution.



c. Write the balanced redox reaction that occurs in basic solution.



# Chapter 9 Stoichiometry Section 1 Answers

**Thomas Evangelist**



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