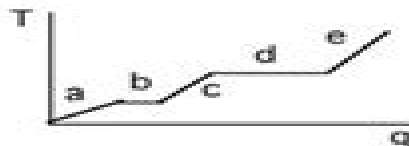


6. The highest vapor pressure is observed for which of the following liquid/temperature combinations?
- $C_3H_8$  at 270 K
  - $C_2H_6$  at 300 K
  - $C_3H_{12}$  at 300 K
  - $HOC_6H_5OH$  at 300 K
  - $HOC_6H_5OH$  at 270 K



7. Which regions on the on the heating curve shown (Temperature versus heat, "q") corresponds to
- a pure gas increasing in temperature **e**
  - a pure liquid increasing in temperature **c**
  - a solid increasing in temperature **a**
  - a solid melting **b**
  - a liquid boiling **d**

8. On the above heating curve, which phases are present:

- In region "a" **solid**
  - In region "b" **solid/liquid**
  - In region "c" **liquid**
  - In region "d" **liquid/gas**
  - In region "e" **gas**
9. A solid is soft, and has a low melting point (below 100°C). Its solid, molten and aqueous forms are all nonconductors. Classify the solid.
10. Indicate the type of crystalline solid each of the following would form upon solidification. Identify the type of particles found at the lattice points and the types of attractive forces.

	Type of crystalline solid	Type of particles at lattice points	IMF
(a)	<b>molecular solid</b>	<b>atoms</b>	<b>molecular bond</b>
(b)	<b>molecular</b>	<b>molecules</b>	<b>dipole-dipole</b>
(c)	<b>molecular</b>	<b>molecules</b>	<b>LDF</b>
graphite	<b>covalent network</b>	<b>atoms</b>	<b>covalent bond</b>

# Chapter Section 12 1 Liquids And Solids Answer Key

**R Barnett**



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