

# LAB 17: LE CHATELIER'S PRINCIPLE

- This lab does NOT require a lab report.
- You will work with a partner for this lab.

## **I** Introduction

LeChatelier's principle states that: If an equilibrium system is subjected to a stress, the system will react to remove the stress. To remove a stress, a system can only do one of two things: form more products and use up reactants (shift right), OR reverse the reaction and form more reactants, using up products (shift left). In this experiment, you will form several equilibrium systems at different stations. Then, by putting different stresses on the systems, you will observe how equilibrium systems react to a stress.

## **!** Warnings!

- This lab uses concentrated (12 M) HCl, hydrochloric acid. This is EXTREMELY dangerous. It is extremely corrosive to skin and eyes, so wear your safety goggles and apron. Its fumes are choking and poisonous, so use it only in the fume hood.
- This lab uses cobalt chloride dissolved in ethanol. Cobalt chloride is mildly toxic; do not ingest and clean up spills immediately. Wear your safety goggles. Ethanol is flammable, so keep away from open flames.

## **🧪** Procedure

*Experiment 1: Equilibrium in a Saturated Solution.*

You will investigate the equilibrium in a saturated sodium chloride solution:



### Directions

- 1) Set up your lab notebook.
- 2) Fill a test tube  $\frac{3}{4}$  full of the saturated NaCl solution (you can see the excess NaCl(s) lying on the beaker bottom to verify its saturation). Be careful not to transfer any solid NaCl into the test tube. This is your initial equilibrium state.
- 3) To this saturated (and solid-free) solution of NaCl, add some Cl<sup>-</sup> ions in the form of concentrated HCl.

	Initial Equilibrium Condition	Stress	Predicted Shift	Observations	Observed Shift
Experiment 1	Saturated NaCl(aq) solution	Add concentrated HCl(aq)			

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**JS Bruner**



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## **Table of Contents Chemfax Applications Of Le Chatelier Lab Answers**

1. Understanding the eBook Chemfax Applications Of Le Chatelier Lab Answers
  - The Rise of Digital Reading Chemfax Applications Of Le Chatelier Lab Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemfax Applications Of Le Chatelier Lab Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemfax Applications Of Le Chatelier Lab Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemfax Applications Of Le Chatelier Lab Answers
  - Personalized Recommendations

- Chemfax Applications Of Le Chatelier Lab Answers User Reviews and Ratings
- Chemfax Applications Of Le Chatelier Lab Answers and Bestseller Lists
- 5. Accessing Chemfax Applications Of Le Chatelier Lab Answers Free and Paid eBooks
  - Chemfax Applications Of Le Chatelier Lab Answers Public Domain eBooks
  - Chemfax Applications Of Le Chatelier Lab Answers eBook Subscription Services
  - Chemfax Applications Of Le Chatelier Lab Answers Budget-Friendly Options
- 6. Navigating Chemfax Applications Of Le Chatelier Lab Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemfax Applications Of Le Chatelier Lab Answers Compatibility with Devices
  - Chemfax Applications Of Le Chatelier Lab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemfax Applications Of Le Chatelier Lab Answers
  - Highlighting and Note-Taking Chemfax Applications Of Le Chatelier Lab Answers
  - Interactive Elements Chemfax Applications Of Le Chatelier Lab Answers
- 8. Staying Engaged with Chemfax Applications Of Le Chatelier Lab Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemfax Applications Of Le Chatelier Lab Answers
- 9. Balancing eBooks and Physical Books Chemfax Applications Of Le Chatelier Lab Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemfax Applications Of Le Chatelier Lab Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemfax Applications Of Le Chatelier Lab Answers
  - Setting Reading Goals Chemfax Applications Of Le Chatelier Lab Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemfax Applications Of Le Chatelier Lab Answers
  - Fact-Checking eBook Content of Chemfax Applications Of Le Chatelier Lab Answers

- Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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