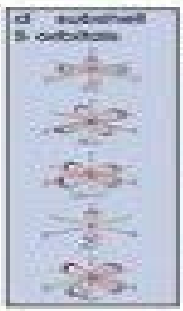


# CHEMISTRY

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## PERIODIC TABLE OF THE ELEMENTS

**Atomic Structure Labels:**  
 H: Atomic Number, Atomic Symbol, Atomic Weight



### ATOMIC STRUCTURE

Atomic Number,  $Z$ , of an element is the number of protons in the nucleus. For a neutral atom,  $Z = +$  the number of electrons. The mass number,  $A$ , is the sum of protons and neutrons in the nucleus. Isotopes of an element have the same  $Z$ , different  $A$ . The atomic mass is the weighted average of isotopes. The atomic weight is the weighted average of isotopes for a given element.

### ATOMIC QUANTUM NUMBERS AND ORBITALS

**Principal Quantum Number ( $n$ ):**  $n = 1, 2, 3, \dots$   
**Angular Momentum Quantum Number ( $l$ ):**  $l = 0, 1, 2, \dots, n-1$   
**Magnetic Quantum Number ( $m_l$ ):**  $m_l = -l, -(l-1), \dots, 0, \dots, (l-1), l$   
**Spin Quantum Number ( $m_s$ ):**  $m_s = +\frac{1}{2}, -\frac{1}{2}$

### NUCLEAR CHEMISTRY

Nuclear reactions alter the number, particles, mass, and energy of atoms. The atomic number,  $Z$ , is the number of protons in the nucleus. The mass number,  $A$ , is the sum of protons and neutrons. The number of neutrons is  $A - Z$ . The number of electrons is  $Z$ . The number of protons is  $Z$ . The number of neutrons is  $A - Z$ . The number of electrons is  $Z$ . The number of protons is  $Z$ . The number of neutrons is  $A - Z$ . The number of electrons is  $Z$ .

### MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the behavior of electrons in atoms. The number of orbitals in a shell is  $n^2$ . The number of orbitals in a subshell is  $2l + 1$ . The number of orbitals in a subshell is  $2l + 1$ . The number of orbitals in a subshell is  $2l + 1$ . The number of orbitals in a subshell is  $2l + 1$ .

n	l	m <sub>l</sub>	m <sub>s</sub>	Orbitals
1	0	0	$+\frac{1}{2}, -\frac{1}{2}$	1 orbital
2	0	0	$+\frac{1}{2}, -\frac{1}{2}$	1 orbital
2	1	0, $\pm 1$	$+\frac{1}{2}, -\frac{1}{2}$	3 orbitals
3	0	0	$+\frac{1}{2}, -\frac{1}{2}$	1 orbital
3	1	0, $\pm 1$	$+\frac{1}{2}, -\frac{1}{2}$	3 orbitals
3	2	0, $\pm 1, \pm 2$	$+\frac{1}{2}, -\frac{1}{2}$	5 orbitals
4	0	0	$+\frac{1}{2}, -\frac{1}{2}$	1 orbital
4	1	0, $\pm 1$	$+\frac{1}{2}, -\frac{1}{2}$	3 orbitals
4	2	0, $\pm 1, \pm 2$	$+\frac{1}{2}, -\frac{1}{2}$	5 orbitals
4	3	0, $\pm 1, \pm 2, \pm 3$	$+\frac{1}{2}, -\frac{1}{2}$	7 orbitals

**Atomic Configuration:** The arrangement of electrons in an atom is described by the atomic configuration. The atomic configuration is the arrangement of electrons in an atom. The atomic configuration is the arrangement of electrons in an atom. The atomic configuration is the arrangement of electrons in an atom.



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**Michael Rajnik**



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