

**Introduction**

This course is a general chemistry course for students who are interested in pursuing a degree in a science field. The course covers the basic principles of chemistry, including atomic structure, chemical bonding, and chemical reactions. The course is designed to provide students with a solid foundation in chemistry, which is essential for many careers in science and technology.

**Prerequisites:** Students must have completed high school chemistry or a college-level chemistry course with a grade of C or better. Students who do not meet these prerequisites should consult with an advisor for more information.

**Part 1: Atomic Structure**



1. Define the terms atom, element, and compound. Explain the difference between an atom and a molecule. Describe the structure of an atom, including the nucleus and the electron cloud.
2. Explain the periodic table of elements. Describe the trends in atomic size, ionization energy, and electronegativity.
3. Calculate the atomic mass of an element from its isotopic composition.



4. Calculate the mass percent of an element in a compound.
5. Calculate the empirical formula of a compound from its mass percent composition.

Element	Atomic Mass	Number of Atoms	Mass of Element	Mass of Compound	Mass Percent
H	1.008	10	10.08	100.00	10.08%
O	16.00	10	160.00	100.00	160.00%
C	12.01	1	12.01	100.00	12.01%
N	14.01	1	14.01	100.00	14.01%

6. Calculate the molecular formula of a compound from its empirical formula and molar mass.
7. Calculate the percent yield of a reaction.

**Part 2: Chemical Bonding**



8. Explain the difference between ionic and covalent bonding. Describe the formation of an ionic bond and a covalent bond.
9. Calculate the lattice energy of an ionic compound.
10. Calculate the bond energy of a covalent bond.

$$\Delta H_{\text{rxn}} = \sum \Delta H_{\text{f}}(\text{products}) - \sum \Delta H_{\text{f}}(\text{reactants})$$

11. Calculate the enthalpy change of a reaction.

Reaction	Enthalpy Change ( $\Delta H_{\text{rxn}}$ )	Equilibrium Constant ( $K_{\text{eq}}$ )	Free Energy Change ( $\Delta G_{\text{rxn}}$ )	Entropy Change ( $\Delta S_{\text{rxn}}$ )
1. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$	-184.6 kJ/mol	> 1	< 0	> 0
2. $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$	-92.4 kJ/mol	> 1	< 0	< 0
3. $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$	-1411 kJ/mol	> 1	< 0	< 0
4. $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$	178.3 kJ/mol	< 1	> 0	> 0

Page 1 of 10 pages. Please refer to the back of this page for more information.

# Basic Stoichiometry Phet Lab Homework Exercises Answers

**IM Harris**



**Basic Stoichiometry Phet Lab Homework Exercises Answers:**

## **Basic Stoichiometry Phet Lab Homework Exercises Answers** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the power of words has be more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such could be the essence of the book **Basic Stoichiometry Phet Lab Homework Exercises Answers**, a literary masterpiece that delves deep to the significance of words and their effect on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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### **Table of Contents Basic Stoichiometry Phet Lab Homework Exercises Answers**

1. Understanding the eBook Basic Stoichiometry Phet Lab Homework Exercises Answers
  - The Rise of Digital Reading Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Basic Stoichiometry Phet Lab Homework Exercises Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Personalized Recommendations
  - Basic Stoichiometry Phet Lab Homework Exercises Answers User Reviews and Ratings

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---

- Basic Stoichiometry Phet Lab Homework Exercises Answers and Bestseller Lists
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  - Basic Stoichiometry Phet Lab Homework Exercises Answers Public Domain eBooks
  - Basic Stoichiometry Phet Lab Homework Exercises Answers eBook Subscription Services
  - Basic Stoichiometry Phet Lab Homework Exercises Answers Budget-Friendly Options
- 6. Navigating Basic Stoichiometry Phet Lab Homework Exercises Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Basic Stoichiometry Phet Lab Homework Exercises Answers Compatibility with Devices
  - Basic Stoichiometry Phet Lab Homework Exercises Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Highlighting and Note-Taking Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Interactive Elements Basic Stoichiometry Phet Lab Homework Exercises Answers
- 8. Staying Engaged with Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Basic Stoichiometry Phet Lab Homework Exercises Answers
- 9. Balancing eBooks and Physical Books Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Basic Stoichiometry Phet Lab Homework Exercises Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Setting Reading Goals Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Fact-Checking eBook Content of Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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