

17. SOLUBILITY PRODUCT CONSTANT

GOALS

1. Become aware of solubility product constants
2. Estimate a solubility constant
3. Observe that a solubility constant is a special case of an equilibrium constant

BACKGROUND

Ionic compounds, such as NaCl , dissociate to form ions when they dissolve in water. An example is sodium chloride:



Where aq is a solution and s is a solid.

Solid NaCl dissociates to become Na^+ and Cl^- ions. The Na^+ cations are attracted to the negative "ends" of water molecules, and the Cl^- anions are attracted to the positive "ends" of water molecules.

In other compounds the attraction between the cations and the anions is stronger and there is little dissociation. These are sparingly or slightly soluble salts. An example is silver chloride, AgCl . A beaker of this solution often has solid AgCl crystals in the bottom. The crystals are in equilibrium with the Ag^+ and Cl^- ions that are in solution:



In this reaction the rate of AgCl dissolving is equal to the rate of Ag^+ and Cl^- ions combining to form solid AgCl . In other words, the reaction is going both ways at the same time and at the same rate. It is said to be in equilibrium. The equilibrium constant is:



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