Stoichiometry III – Theoretical Yield

Teacher

If we know what chemical reaction is occurring, we can use mole ratios to determine the limiting and excess reactants. We can do this by using the starting masses of each reactant to see how much of a product could be produced if all of the reactant was consumed. We then compare the number of moles produced by each reactant, and the smallest value is the **theoretical yield** – the amount of a product produced if all of the limiting reactant is consumed.

Practice Problems:

1. Given the following reaction: $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

   a) Determine the limiting and excess reagents if you started with 14.8 g of C_3H_8 and 33.44 g of O_2.

   $$\frac{14.8_g\ C_3H_8}{} \mid \frac{1\ mol\ C_3H_8}{44.09_g\ C_3H_8} \mid \frac{4\ mol\ H_2O}{1\ mol\ C_3H_8} \mid \frac{18.02_g\ H_2O}{1\ mol\ H_2O} = 24.2_g\ H_2O$$

   $$\frac{33.44_g\ O_2}{} \mid \frac{1\ mol\ O_2}{32.00_g\ O_2} \mid \frac{4\ mol\ H_2O}{5\ mol\ O_2} \mid \frac{18.02_g\ H_2O}{1\ mol\ H_2O} = 15.1_g\ H_2O$$

   Limiting: O_2
   Excess: C_3H_8

   b) What is the theoretical yield of H_2O. $15.1_g\ H_2O$

2. Given the following reaction: $Al_2(SO_3)_3 + 6NaOH \rightarrow 3Na_2SO_3 + 2Al(OH)_3$

   a) Determine the limiting and excess reagents if 8.1 g of $Al_2(SO_3)_3$ is reacted with 8.1 g of NaOH

   $$\frac{8.1_g\ Al_2(SO_3)_3}{} \mid \frac{1\ mol\ Al_2(SO_3)_3}{294.14_g\ Al_2(SO_3)_3} \mid \frac{2\ mol\ Al(OH)_3}{1\ mol\ Al_2(SO_3)_3} \mid \frac{78.00_g\ Al(OH)_3}{1\ mol\ Al(OH)_3} = 4.3_g\ Al(OH)_3$$

   $$\frac{8.1_g\ NaOH}{} \mid \frac{1\ mol\ NaOH}{40.00_g\ NaOH} \mid \frac{2\ mol\ Al(OH)_3}{6\ mol\ NaOH} \mid \frac{78.00_g\ Al(OH)_3}{1\ mol\ Al(OH)_3} = 5.3_g\ Al(OH)_3$$

   Limiting: $Al_2(SO_3)_3$
   Excess: NaOH

   b) What is the theoretical yield of aluminum hydroxide. $4.3_g\ Al(OH)_3$

3. Given the following reaction: $4Al_2O_3 + 9Fe \rightarrow 3Fe_3O_4 + 8Al$

   a) Determine the limiting and excess reagents if 25.4 g of Al_2O_3 is reacted with 10.2 g of Fe.

   $$\frac{25.4_g\ Al_2O_3}{} \mid \frac{1\ mol\ Al_2O_3}{101.96_g\ Al_2O_3} \mid \frac{3\ mol\ Fe_3O_4}{4\ mol\ Al_2O_3} \mid \frac{231.55_g\ Fe_3O_4}{1\ mol\ Fe_3O_4} = 43.3_g\ Fe_3O_4$$

   $$\frac{10.2_g\ Fe}{} \mid \frac{1\ mol\ Fe}{55.85_g\ Fe} \mid \frac{3\ mol\ Fe_3O_4}{9\ mol\ Fe} \mid \frac{231.55_g\ Fe_3O_4}{1\ mol\ Fe_3O_4} = 14.1_g\ Fe_3O_4$$

   Limiting: Fe
   Excess: Al_2O_3

   b) What is the theoretical yield of iron(II,III) oxide (aka magnetite).

   $14.1_g\ Fe_3O_4$

Chapter 3 Stoichiometry Exercises Answers

R Barnett

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