

Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

Chapter 6 Chemistry Chemical Bonding

Ibrahim Sikder



Chapter 6 Chemistry Chemical Bonding:

A-level Chemistry E. N. Ramsden, 2000 Each topic is treated from the beginning without assuming prior knowledge Each chapter starts with an opening section covering an application These help students to understand the relevance of the topic they are motivational and they make the text more accessible to the majority of students Concept Maps have been added which together with Summaries throughout aid understanding of main ideas and connections between topics Margin points highlight key points making the text more accessible for learning and revision Checkpoints in each chapter test students understanding and support their private study Topological Approaches to the Chemical Bond Ángel Martín Pendás, Julia Contreras-García, 2023-02-22 This graduate textbook provides comprehensive information on topological analysis in real space of the electronic structure Application of the topological tools is becoming routine for understanding the outcome of quantum chemical calculations This title thoroughly reviews a selection of currently available topological tools their use and spectrum of applications and provides graduate students and researchers with information not easily obtained from the available textbooks The book is accompanied by worked examples exercises and solutions and is a great tool for any quantum chemistry or computational chemistry course at the graduate and advanced undergraduate levels The Chemical Bond Ahmed Zewail, 1992-05-14 This inspired book by some of the most influential scientists of our time including six Nobel laureates chronicles our emerging understanding of the chemical bond through the last nine decades and into the future From Pauling's early structural work using x ray and electron diffraction to Zewail's femtosecond lasers that probe molecular dynamics in real time from Crick's molecular biology to Rich's molecular recognition this book explores a rich tradition of scientific heritage and accomplishment The perspectives given by Pauling Perutz Rich Crick Porter Polanyi Herschbach Zewail and Bernstein celebrate major scientific achievements in chemistry and biology with the chemical bond playing a fundamental role In a unique presentation that also provides some lively insights into the very nature of scientific thought and discovery The Chemical Bond Structure and Dynamics will be of general interest to scientists science historians and the scientifically inclined populous The Chemical Bond Tadamasu Shida, 2013-03-09 Providing the quantum mechanical foundations of chemical bonding this unique textbook emphasizes key concepts such as superposition degeneracy of states and the role of the electron spin These quantum mechanical notions are usually oversimplified or meticulously circumvented in other books to the frustration of serious readers who want to understand for example why two protons can be stably bound with only one electron to make the simplest molecule H_2 An initial concise and compact presentation of the rudiments of quantum mechanics enables readers to progress through the book with a firm grounding Experimental examples are included to illustrate how the abstract concepts are manifest in real systems **Chemistry Textbook for College and University USA** Ibrahim Sikder, 2023-06-04 *Chemistry Textbook USA* *Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science*, 2003-11 Prentice Hall Physical Science Concepts in Action helps

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