

# 8

## Acids, bases and salts

### Key objectives

By the end of this section, you should be able to:

- describe the characteristic reactions of acids
  - state that bases are oxides or hydroxides of metals
  - state that alkalis are soluble bases
  - describe the characteristic reactions of bases
  - state that aqueous solutions of acids contain  $H^+$  ions and aqueous solutions of alkalis contain  $OH^-$  ions
  - describe how to use pH, as measured with universal indicator paper, to compare hydrogen ion concentration, neutrality, relative acidity and alkalinity
  - describe the neutralisation reaction
- define acids and bases in terms of proton transfer
  - explain the difference between weak and strong acids in terms of dissociation
- classify oxides as acidic, basic or amphoteric, with examples
  - describe the general solubility rules for salts
  - describe the preparation of soluble salts by the reaction of an acid with:
    - an alkali
    - excess metal, insoluble base or carbonate
- describe the preparation of insoluble salts by precipitation
- define a hydrated substance and an anhydrous substance
- define the term water of crystallisation in crystals, including  $CuSO_4 \cdot 5H_2O$  and  $CoCl_2 \cdot 6H_2O$

### Key terms

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Term	Definition
<b>Acid</b>	A substance which dissolves in water, producing $H^+(aq)$ ions as the only positive ion. <b>A proton (<math>H^+</math>) donor.</b>
<b>Alkali</b>	A soluble base which produces $OH^-(aq)$ ions in water.
<b>Anhydrous salt</b>	A salt which has had its water of crystallisation removed.
<b>Base</b>	A substance which neutralises an acid, producing a salt and water as the only products. Bases are oxides or hydroxides of metals. (Ammonia is also a base.) <b>A proton (<math>H^+</math>) acceptor.</b>
<b>Indicator</b>	A substance that shows whether a substance is acidic or alkaline by changing colour.
<b>Neutralisation</b>	The process in which an acid reacts with a base to form water.
<b>pH scale</b>	A scale running from 0 to 14 used to express the acidity or alkalinity of a substance.
<b>Saturated solution</b>	A solution containing the maximum amount of dissolved solute in the solvent at a given temperature.
<b>Water of crystallisation</b>	Water incorporated into the structure of a substance as it crystallises, for example in copper(II) sulfate pentahydrate ( $CuSO_4 \cdot 5H_2O$ ).

### 8.1 Acids and alkalis

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#### Acids

**Acids** are substances that produce  $H^+$  ions when they are dissolved in water.

Acids are defined as proton ( $H^+$ ) donors.

# Chapter 8 Acids Bases And Ph Quia

**JR Anderson**



## **Chapter 8 Acids Bases And Ph Quia:**

## Reviewing **Chapter 8 Acids Bases And Ph Quia**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Chapter 8 Acids Bases And Ph Quia**," an enthralling opus penned by a very acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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