

CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron III sulfate, the roman numeral tells us **C**
 - How many atoms of Fe are in one formula unit.
 - How many sulfate ions can be attached to the iron atom
 - The charge on each Fe ion
 - The total positive charge of the formula unit
- The result of changing a subscript in a correctly written chemical formula is to **c**
 - Change the number of moles represented in the formula
 - Change the charges on the other ions in the compound
 - Change the formula so that it no longer represents the compound
 - Have no effect on the formula
- Derive a generalization for when an acid name will end in the suffix **-ic** or **-ous**
If ending is -ic, the anion ended in -ate; if ending is -ous, the anion ended in -ite
- Derive a generalization for when an acid name will begin with the prefix **hydro-**.
If prefix is hydro-, the anion ended in -ide. This is most likely a monatomic ion.
- Fill in the blanks in the table below

Compound Name	Formula
Aluminum sulfide	Al_2S_3
Aluminum sulfate	$Al_2(SO_4)_3$
Lead II chloride	$PbCl_2$
Ammonium phosphite	$(NH_4)_2PO_3$
Hydroiodic acid	HCl
Dinitrogen pentoxide	N_2O_5
Copper II carbonate	$CuCO_3$
Sodium sulfide	Na_2S
Ammonium phosphate	$(NH_4)_3PO_4$
Tin IV sulfite	$Sn(SO_3)_2$
Nitrous acid	HNO_2
Magnesium perchlorate	$Mg(ClO_4)_2$
Iron II nitrate	$Fe(NO_3)_2$
Iron III nitrate	$Fe(NO_3)_3$
Cobalt II oxide	CoO
Sulfur trioxide	SO_3
Boron pentafluoride	BF_5
Tetraphosphorus octasulfide	P_4S_8
Tetraiodide nonoxide	I_4O_9

- Name the following compound and identify as molecular, ionic or acid
 - $Ca(HCO_3)_2$ **calcium bicarbonate - I**
 - $Fe(ClO_2)_3$ **Iron III chlorite - I**
 - $Fe(ClO_4)_2$ **Iron II chlorate - I**
 - N_2O_5 **dinitrogen pentoxide - M**

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