

**CHEM 101 Quiz #6 Chapter 6 ANSWERS****1. (4 points)**

The specific heat (s) of iron (Fe) is 0.40 J/(g·K). Calculate the heat effect  $q$  when 20 g of iron is heated from 100 °C to 200 °C. Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \text{ in } \Delta K = \Delta ^\circ C \quad q = 20 \times 0.40 \times (200 - 100) = +800 \text{ J}$$

You must include the units and sign in the answer!

**2. (4 points)**

The Molar Mass of Mg = 24.0 g/mol.

Calculate the enthalpy change when 2.40 g Mg(s) reacts with excess oxygen. Include the sign!

$$2.40 \text{ g Mg} = 2.40/24.0 = 0.1 \text{ mol} \rightarrow \Delta H = 0.1(\text{mol}) \times (-600)(\text{kJ/mol}) = -60 \text{ kJ}$$

Again, your answer must include both the sign and the units.

**3. (2 points)** Which one of the following four reactions is a *formation reaction*? Circle the correct answer.



*Look up the definition of a formation reaction in Chang or in your lecture notes: the compound is formed from its elements, and the elements must be as they naturally occur so e.g.  $\text{Cl}_2$  not  $\text{Cl}$  and  $\text{H}_2$  not  $\text{H}$ !*

**CHEM 101 Section L52 Quiz #6****1. (4 points)**

The specific heat (s) of copper (Cu) is 0.38 J/(g·K). Calculate the heat effect  $q$  when 10 g of copper is heated from 100 °C to 300 °C. Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \text{ in } \Delta K = \Delta ^\circ C \quad q = 10 \times 0.38 \times (300 - 100) = +760 \text{ J}$$

You must include the units and sign in the answer!

**2. (4 points)**

The Molar Mass of Ca = 40.0 g/mol.

Calculate the enthalpy change  $\Delta H$  when 8.0 g Ca(s) reacts with excess oxygen. Include the sign!!

$$8.0 \text{ g Ca} = 8.0/40 = 0.2 \text{ mol} \rightarrow \Delta H = 0.2(\text{mol}) \times (-650)(\text{kJ/mol}) = -130 \text{ kJ}$$

Again, your answer must include both the sign and the units.

**3. (2 points)** Which one of the following reactions is a *formation reaction*? Circle the correct answer.



# Chem Fax Lab 6 Answers

**LL Leslie**



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