

Chapter Test II

A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description.

Column I

- _____ 1. An unbalanced chemical equation for a reaction that obeys the law of conservation of matter.
- _____ 2. The number of protons in an atom of an element. It is equal to the number of electrons.
- _____ 3. The number of electrons in an atom of an element.
- _____ 4. The ratio of the actual weight of an element to the standard weight of 1 gram.
- _____ 5. The number of atoms of an element in a molecule of a compound.
- _____ 6. The number of atoms of an element in a molecule of an element.

Column II

- a. atomic weight
- b. limiting reagent
- c. molecular weight
- d. stoichiometry
- e. atomic number
- f. atomic weight
- g. atomic weight

B. Multiple Choice

Choose the correct answer for each question.

- _____ 1. How many molecules of H_2O are formed when 100 g of H_2 reacts with 100 g of O_2 ?

<ol style="list-style-type: none"> a. 1000 molecules b. 10000 molecules c. 100000 molecules d. 1000000 molecules 	<ol style="list-style-type: none"> e. 10000000 molecules f. 100000000 molecules g. 1000000000 molecules h. 10000000000 molecules
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- _____ 2. Which of the following is not a chemical change?
 - a. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
 - b. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$
 - c. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$
 - d. $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2 + \text{H}_2\text{O}$
- _____ 3. How many grams of H_2 are required to reduce 100 g of Fe_2O_3 to Fe ?

<ol style="list-style-type: none"> a. 100 g b. 1000 g c. 10000 g d. 100000 g 	<ol style="list-style-type: none"> e. 1000000 g f. 10000000 g g. 100000000 g h. 1000000000 g
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Chemistry Chapter 12 Stoichiometry Test

John Moore, Richard H. Langley



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tables and a glossary of chemistry terms

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