

## Chapter 9 Stoichiometry Chapter Review Answers

- Mole Ratios
  - Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
  - They are formed from the coefficients found in a balanced chemical equation.
- Mole Ratios as conversion factors
  - $\frac{2 \text{ mol Ca}}{1 \text{ mol O}_2} \cdot \frac{1 \text{ mol O}_2}{2 \text{ mol Ca}} = \frac{2 \text{ mol CaO}}{2 \text{ mol Ca}} = \frac{2 \text{ mol Ca}}{2 \text{ mol CaO}} = \frac{1 \text{ mol O}_2}{2 \text{ mol CaO}} = \frac{2 \text{ mol CaO}}{1 \text{ mol O}_2}$ . Remember the ratios can be reduced as well!!!
  - $\frac{1 \text{ mol Mg}}{2 \text{ mol HF}} \cdot \frac{2 \text{ mol HF}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol Mg}} = \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} = \frac{1 \text{ mol H}_2}{2 \text{ mol HF}} = \frac{1 \text{ mol Mg}}{1 \text{ mol MgF}_2} = \frac{1 \text{ mol H}_2}{2 \text{ mol HF}} = \frac{1 \text{ mol HF}}{1 \text{ mol Mg}} = \frac{1 \text{ mol MgF}_2}{2 \text{ mol HF}} = \frac{2 \text{ mol HF}}{1 \text{ mol MgF}_2} = \frac{1 \text{ mol H}_2}{2 \text{ mol HF}} = \frac{1 \text{ mol MgF}_2}{1 \text{ mol H}_2} = \frac{1 \text{ mol H}_2}{1 \text{ mol MgF}_2}$
- $\frac{103.99 \text{ g Na}_2\text{CO}_3}{1 \text{ mol Na}_2\text{CO}_3} \cdot \frac{1 \text{ mol Na}_2\text{CO}_3}{103.99 \text{ g Na}_2\text{CO}_3} = \frac{74.10 \text{ g Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} = \frac{1 \text{ mol Ca(OH)}_2}{74.10 \text{ g Ca(OH)}_2} = \frac{40.00 \text{ g NaOH}}{1 \text{ mol NaOH}} = \frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} = \frac{100.09 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} = \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3}$
- Molar Mass
  - Definition: the mass of one mole of a substance
  - Molar mass is used as a conversion factor in stoichiometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
- Hydrogen + Oxygen  $\rightarrow$  Water
  - 5.0 mol H<sub>2</sub>
  - 2.5 mol O<sub>2</sub>
- Ethane combustion reaction
  - 15.8 mol O<sub>2</sub>
  - 9.00 mol CO<sub>2</sub> & 13.5 mol H<sub>2</sub>O
- 575 g Na & 886 g Cl<sub>2</sub>
- Iron production reaction
  - 75.1 mol CO
  - 50.1 mol Fe & 75.1 mol CO<sub>2</sub>
- 87,390 g CO or  $8.739 \times 10^4$  g CO & 12,610 g H<sub>2</sub> or  $1.261 \times 10^4$  g H<sub>2</sub>
- Lightning reaction of N<sub>2</sub> & O<sub>2</sub>
  - 1,100 g NO<sub>2</sub> or  $1.10 \times 10^3$  g NO<sub>2</sub>
  720. g NO or  $7.20 \times 10^2$  g NO
- CO<sub>2</sub> removal in spacecraft
  - 42.04 mol NaOH
  - 21.02 mol Na<sub>2</sub>CO<sub>3</sub> & 21.02 mol H<sub>2</sub>O
- Double replacement AgNO<sub>3</sub> & NaBr
  - 463 g NaBr
  - 845 g AgBr
- Soda-acid fire extinguisher
  - 3.058 mol NaHCO<sub>3</sub>
  - 3.058 mol CO<sub>2</sub>, 1.529 mol Na<sub>2</sub>SO<sub>4</sub>, 3.058 mol H<sub>2</sub>O
- Sulfuric acid-sodium hydroxide neutralization reaction
  - H<sub>2</sub>SO<sub>4</sub> + 2 NaOH  $\rightarrow$  Na<sub>2</sub>SO<sub>4</sub> + 2 H<sub>2</sub>O
  - 37 g H<sub>2</sub>SO<sub>4</sub>
  - 53 g Na<sub>2</sub>SO<sub>4</sub> & 14 g H<sub>2</sub>O
- Copper & silver nitrate reaction
  - 0.0104 mol Cu(NO<sub>3</sub>)<sub>2</sub>
  - 0.0104 mol Cu & 0.0209 mol AgNO<sub>3</sub>
- Aspirin production
  - 13.5 kg C<sub>9</sub>H<sub>8</sub>O<sub>4</sub>
  - 7.66 kg C<sub>7</sub>H<sub>6</sub>O<sub>3</sub>
  - 4.29 L HC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>
- Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
- The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant **limits** the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

# Chemistry Chapter 9 Stoichiometry Answers

**Michael Hoelscher**



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**Answer Book to Accompany Boikess/Edelson, Chemical Principles, Third Edition** Robert S. Boikess, 1985

**Chemistry Calculations for Beginners** John Obimakinde, Samuel Obimakinde, Ebenezer Obimakinde, Fredrick Akinbolade, 2025-05-30 With decades of combined experience as science teachers at both school and undergraduate levels the authors have recognised that one of the greatest challenges faced by students studying chemistry is grasping the complexity of the numerous numerical problems found in most parts of the subject This text is crafted to provide a clear and accessible pathway to overcoming this challenge by assisting students especially novices or those with minimal knowledge of the subject in performing chemistry calculations The content covers fundamental calculations crucial to understanding the principles of chemistry making it an invaluable tool for students aiming to excel in their studies Key features Designed with a student friendly approach including detailed explanation of chemical concepts underlying each type of calculation step by step explanations alternative methods for solving problems numerous practice exercises answers to practice exercises and appendices The book is tailored to suit various curricula ensuring relevance for a diverse audience Encompasses a wide range of calculations offering students a thorough understanding of essential chemistry concepts Serves as an excellent resource for exam preparation and equips students with skills applicable to future scientific endeavours Employs straightforward language to ensure ease of understanding for beginners Uses IUPAC conventions underscoring the universal nature of chemistry *McGraw-Hill's SAT Subject Test: Chemistry, 2ed* Thomas Evangelist, 2009-02-01 We want to help you score high on the SAT Chemistry test We've put all of our proven expertise into McGraw Hill's SAT Subject Test Chemistry to make sure you're fully prepared for this difficult exam With this book you'll get essential skill building techniques and strategies created by a leading high school chemistry teacher You'll also get 5 full length practice tests hundreds of sample questions and all the facts about the current exam With McGraw Hill's SAT Subject Test Chemistry we'll guide you step by step through your preparation program and give you the tools you need to succeed 5 full length sample tests with complete explanations for every question 40 top test items to remember on exam day A step by step review of all topics covered on the exam Teacher recommended tips and strategies to help you raise your score **McGraw-Hill's SAT Subject Test Chemistry, 3rd Edition** Thomas Evangelist, 2011-02-03 Expert guidance on the Chemistry exam Many colleges and universities require you to take one or more SAT II Subject Tests to demonstrate your mastery of specific high school subjects McGraw Hill's SAT Subject Test Chemistry is written by experts in the field and gives you the guidance you need

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web the amino acid solutions were at approximately pH 6 prior to the addition of formaldehyde the end point of the titration was at about pH 9 no correction was made for the volume of standard base required to bring the aqueous formaldehyde solution to the end point pH of the amino acid titration and the concentration of the formaldehyde

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