

## CHEMISTRY PRACTICAL - SALT ANALYSIS

**NOTE:**

1. Exp. 1 and 2 for your reference, no need to write in your notebook.
2. Exp. 3 to 10 are to be written on the given table.
3. Chemical Equations are to be learned by the students thoroughly.
4. Start each Experiment from new page.

### EXPERIMENT – 1 Classification of Anions

Group	Group Reagent	Observation	Inference
A	Dilute $H_2SO_4$	a) Colourless, odourless gas with brisk effervescence ( $CO_2$ ) which turns lime water milky. b) Colourless gas with rotten egg like smell ( $H_2S$ ) which turns lead acetate paper black. c) Colourless gas with smell of burning sulphur ( $SO_2$ ) which turns acidified dichromate paper green. d) Brown coloured gas ( $NO_2$ ) which turns ferrous sulphate solution black or brown. e) Colourless gas with vinegar like smell.	$CO_3^{2-}$ (Carbonate) $S^{2-}$ (sulphide) $SO_3^{2-}$ (Sulphite) $NO_2^-$ (Nitrite) $CH_3COO^-$ (Acetate)
B	Conc. $H_2SO_4$	a) Colourless pungent smelling gas ( $HCl$ ) which gives white dense fumes with glass rod dipped in $NH_4OH$ . b) Violet coloured vapours ( $I_2$ ) which turns starch paper blue. c) Reddish brown gas ( $NO_2$ ) having pungent smell (On adding copper turning, fumes becomes intense) d) Brown colour gas with pungent smell ( $Br_2$ ) which turns starch paper yellow. e) Colourless, odourless gas with brisk effervescence ( $CO + CO_2$ ) which turns lime water milky and burns on the mouth of test tube with blue flame.	$Cl^-$ (Chloride) $I^-$ (Iodide) $NO_3^-$ (Nitrate) $Br^-$ (Bromide) $C_2O_4^{2-}$ (Oxalate)
C	$BaCl_2$	White ppt. of $BaSO_4$ is formed.	$SO_4^{2-}$ (sulphate)
D	Ammonium molybdate $3(NH_4)_2MoO_4$	Cannary yellow ppt. of phospho ammonium molybdate $(NH_4)_3PO_4 \cdot 12 MoO_3 \cdot 6H_2O$	$PO_4^{3-}$ (phosphate)

### EXPERIMENT – 2

#### Classification of Cations

Group	Group Reagent	Radical	PPt/Smell	Colour
Zero	$NaOH$	$NH_4^+$	Smell of $NH_3$	-
I	Dil. $HCl$	$Pb^{2+}$	$PbCl_2$	White
II	$H_2S$ gas in acidic Medium	$Pb^{2+}$ $As^{3+}$ $Cu^{2+}$ $Cd^{2+}$	$PbS$ $As_2S_3$ $CuS$ $CdS$	Black Yellow Black Yellow
III	$NH_4Cl$ (s) in presence of $NH_4OH$	$Fe^{3+}$ $Fe^{2+}$ $Al^{3+}$	$Fe(OH)_3$ $Fe(OH)_2$ $Al(OH)_3$	Light green Reddish brown Gelatinous white
IV	$H_2S$ gas in basic medium	$Ni^{2+}$	$NiS$	Black

# Class Xii Chemistry Practical Salt Analysis

**JG Myers**



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