

Solutions

Theme 2

Exercises

1. Define the term solution. How many types of solutions are formed? Write briefly about each type with an example.

Ans. A solution is a homogeneous mixture of two or more pure substances. Solute and solvent are two components of a solution.

Types of solutions: On the basis of different states of solute and solvent solutions are of three types mainly, which are further categorized into three.

Type of Solution	Solute	Solvent	Common Examples
Gaseous Solutions	Gas	Gas	Mixture of oxygen and nitrogen gases
	Liquid	Gas	Chloroform mixed with nitrogen gas
	Solid	Gas	Camphor in nitrogen gas
Liquid Solutions	Gas	Liquid	Oxygen dissolved in water
	Liquid	Liquid	Ethanol dissolved in water
	Solid	Liquid	Glucose dissolved in water
Solid Solutions	Gas	Solid	Solution of hydrogen in palladium
	Liquid	Solid	Amalgam of mercury with sodium
	Solid	Solid	Copper dissolved in gold

2. Give an example of a solid solution in which the solute is a gas.

Ans. In case a solid solution is formed between two substances, an interstitial solid solution will be formed. For example: a solution of hydrogen in palladium is a solution in which the solute is a gas.

3. Define the following terms:

(i) Mole fraction

(iii) Molarity

(ii) Molality

(iv) Mass percentage

Ans. (i) **Mole fraction:** It is defined as the ratio of the number of moles of the individual component (solute or solvent) to the total number of moles of all components present in the solution. If n_A is the number of moles of solute dissolved in n_B moles of solvent, then Mole fraction of solute

$$(X_A) = \frac{n_A}{n_A + n_B}$$

$$\text{Mole fraction of solvent } (X_B) = \frac{n_B}{n_A + n_B}$$

(ii) **Molality:** It is defined as the number of moles of a solute present in 1000g (1kg) of a solvent.

$$\text{Molality } (m) = \frac{\text{Number of moles of solute}}{\text{Weight of solvent in kg}} = \frac{n}{W}$$

$$\text{Unit} = \text{mol kg}^{-1} \text{ or m(molal)}$$

(iii) **Molarity:** It is defined as the number of moles of solute present in one litre of solution.

$$\text{Molarity } (M) = \frac{\text{Number of moles of solute}}{\text{Volume of solution in litre}} = \frac{n}{V}$$

$$\text{Unit} = \text{mol L}^{-1} \text{ or M(Molar)}$$

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Madsen**



Whispering the Strategies of Language: An Mental Quest through **Chapter 2 Solutions Page 12 Of 28**

In a digitally-driven earth wherever displays reign supreme and immediate communication drowns out the subtleties of language, the profound strategies and mental subtleties hidden within words frequently move unheard. However, situated within the pages of **Chapter 2 Solutions Page 12 Of 28** a fascinating literary prize pulsating with organic feelings, lies an exceptional journey waiting to be undertaken. Published by a skilled wordsmith, that enchanting opus attracts viewers on an introspective trip, delicately unraveling the veiled truths and profound impact resonating within the cloth of each word. Within the emotional depths of this moving evaluation, we shall embark upon a genuine exploration of the book is primary subjects, dissect their charming publishing design, and fail to the powerful resonance it evokes serious within the recesses of readers hearts.

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