

9. Evaluate the experiments that led to the conclusion that electrons are negatively charged particles found in all matter.

The deflection toward positively charged plates demonstrated the negatively charged nature of electrons; the fact that changing the type of electrode or the type of gas used in the cathode ray tube did not affect the ray produced led to the conclusion that electrons are present in all matter.

10. Compare the relative charge and mass of each of the subatomic particles.

Particle	Relative charge	Relative mass
Electron	-1	1/1840
Proton	+1	1
Neutron	0	-1

11. **Calculate** What is the difference expressed in kilograms between the mass of a proton and the mass of an electron?

$$1.673 \times 10^{-24} \text{ g} - 9.11 \times 10^{-28} \text{ g} = 1.672 \times 10^{-24} \text{ g} \times 1 \text{ kg}/1000 \text{ g} = 1.672 \times 10^{-27} \text{ kg}$$

Section 4.3 How Atoms Differ

pages 115–121

Practice Problems

pages 116–121

12. How many protons and electrons are in each atom?
- radon, 86 protons and 86 electrons
 - magnesium, 12 protons and 12 electrons
13. An atom of an element contains 66 electrons. Which element is it?
dysprosium
14. An atom of an element contains 14 protons. Which element is it?
silicon

15. **Challenge** Do the atoms shown in the figure to the right have the same atomic number?



Yes, 9

16. Determine the number of protons, electrons, and neutrons for isotopes b through f, in the table above. Name each isotope, and write its symbol.

	Protons and Electrons	Neutrons	Isotope	Symbol
b.	20	26	calcium-46	${}^{46}_{20}\text{Ca}$
c.	8	9	oxygen-17	${}^{17}_8\text{O}$
d.	26	31	iron-57	${}^{57}_{26}\text{Fe}$
e.	30	34	zinc-64	${}^{64}_{30}\text{Zn}$
f.	80	124	mercury-204	${}^{204}_{80}\text{Hg}$

17. **Challenge** An atom has a mass number of 55. Its number of neutrons is the sum of its atomic number and 5. How many protons, neutrons, and electrons does this atom have? What is the identity of this atom?

25 protons, 25 electrons, 30 neutrons, manganese

18. Boron (B) has two naturally occurring isotopes: boron-10 (abundance = 19.8%, mass = 10.013 amu), boron-11 (abundance = 80.2%, mass = 11.009 amu). Calculate the atomic mass of boron.

$$\text{B-10 } 10.013 \text{ amu} \times 0.198 = 1.98 \text{ amu}$$

$$\text{B-11 } 11.009 \text{ amu} \times 0.802 = 8.83 \text{ amu}$$

$$\text{Atomic mass } 1.98 \text{ amu} + 8.83 \text{ amu} = 10.81 \text{ amu}$$

19. **Challenge** Nitrogen has two naturally occurring isotopes, N-14 and N-15. Its atomic mass is 14.007. Which isotope is more abundant? Explain your answer.

N-14 is more abundant because the atomic mass is closer to 14 than 15.

Chapter 4 Atomic Structure Assessment Answers

K Payea



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Table of Contents Chapter 4 Atomic Structure Assessment Answers

1. Understanding the eBook Chapter 4 Atomic Structure Assessment Answers
 - The Rise of Digital Reading Chapter 4 Atomic Structure Assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 4 Atomic Structure Assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 4 Atomic Structure Assessment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 4 Atomic Structure Assessment Answers

- Personalized Recommendations
 - Chapter 4 Atomic Structure Assessment Answers User Reviews and Ratings
 - Chapter 4 Atomic Structure Assessment Answers and Bestseller Lists
5. Accessing Chapter 4 Atomic Structure Assessment Answers Free and Paid eBooks
 - Chapter 4 Atomic Structure Assessment Answers Public Domain eBooks
 - Chapter 4 Atomic Structure Assessment Answers eBook Subscription Services
 - Chapter 4 Atomic Structure Assessment Answers Budget-Friendly Options
 6. Navigating Chapter 4 Atomic Structure Assessment Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 4 Atomic Structure Assessment Answers Compatibility with Devices
 - Chapter 4 Atomic Structure Assessment Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 4 Atomic Structure Assessment Answers
 - Highlighting and Note-Taking Chapter 4 Atomic Structure Assessment Answers
 - Interactive Elements Chapter 4 Atomic Structure Assessment Answers
 8. Staying Engaged with Chapter 4 Atomic Structure Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 4 Atomic Structure Assessment Answers
 9. Balancing eBooks and Physical Books Chapter 4 Atomic Structure Assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 4 Atomic Structure Assessment Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter 4 Atomic Structure Assessment Answers
 - Setting Reading Goals Chapter 4 Atomic Structure Assessment Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter 4 Atomic Structure Assessment Answers

- Fact-Checking eBook Content of Chapter 4 Atomic Structure Assessment Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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