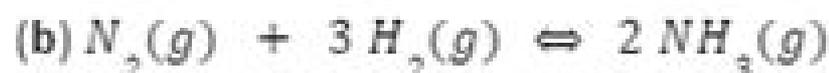


1. Write equilibrium constant expressions,  $K_p$ , for the following reactions.



$$\text{Ans. } K_p = \frac{P_{\text{CO}_2} \times P_{\text{H}_2}}{P_{\text{H}_2\text{O}} \times P_{\text{CO}}}$$



$$\text{Ans. } K_p = \frac{P_{\text{NH}_3}^2}{P_{\text{H}_2}^3 \times P_{\text{N}_2}}$$

2. A mixture containing equal concentrations of methane and steam is passed over a nickel catalyst at 1000 K. The emerging gas has the composition  $[\text{CO}] = 0.1027 \text{ mol dm}^{-3}$ ,  $[\text{H}_2] = 0.3080 \text{ mol dm}^{-3}$ , and  $[\text{CH}_4] = [\text{H}_2\text{O}] = 0.8973 \text{ mol dm}^{-3}$ . Assuming this mixture is at equilibrium, calculate the equilibrium constant  $K_c$  for the reaction

$$\begin{aligned} \text{Ans. } K_c &= \frac{[\text{CO}][\text{H}_2]^3}{[\text{CH}_4][\text{H}_2\text{O}]} = \frac{0.1027 \text{ mol dm}^{-3} \times (0.3080 \text{ mol dm}^{-3})^3}{0.8973 \text{ mol dm}^{-3} \times 0.8973 \text{ mol dm}^{-3}} \\ &= 3.727 \times 10^{-3} \text{ mol}^2 \text{ dm}^{-6} \end{aligned}$$

# Chemical Equilibrium Problems With Solutions

**Albert A Gayle**



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## Unveiling the Magic of Words: A Review of "**Chemical Equilibrium Problems With Solutions**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Chemical Equilibrium Problems With Solutions**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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