

I- 1) X un élément qui appartient au groupe IIA et à la deuxième période, sa couche de valence sera de type ns^2 avec $n = 2$, donc $2s^2$.

Structure électronique de X : $1s^2 2s^2$. Son numéro atomique $Z = 4$.

2) X^{2+} est l'ion hydrogénoïde de X.

3) Constante de Rhydberg relative à X^{2+} est : $K = Z^2 \cdot R_{H1} = 4^2 \cdot 1,09678 \cdot 10^7 = 17,54848 \cdot 10^7 \text{ m}^{-1}$.

4) Longueur d'onde de la radiation capable d'ioniser X^{2+} est :

$$\frac{1}{\lambda_{\text{ion}}} = K \left(\frac{1}{1^2} - \frac{1}{\infty^2} \right) \Rightarrow \lambda_{\text{ion}} = \frac{1}{K} = \frac{1}{17,54848 \cdot 10^7} = 0,05698 \cdot 10^{-7} \text{ m} = 56,98 \text{ \AA}$$

$$5) \Delta E = \frac{hc}{\lambda} = \frac{6,626 \cdot 10^{-34} \cdot 3 \cdot 10^8}{56,9 \cdot 10^{-10}} = 0,348 \cdot 10^{-16} = 348 \cdot 10^{-19} \text{ J} = 2,16 \text{ eV}$$

II- 1) n : nombre quantique principal, il détermine le niveau d'énergie ou la couche associée à l'électron. n ne peut prendre que des valeurs entières : $n(\text{entier}) \geq 1$.

l : nombre quantique secondaire ou azimutal, il détermine la forme générale de l'orbitale c'est-à-dire la région dans laquelle l'électron se déplace (ou configuration spatiale) : $0 \leq l(\text{entier}) \leq n-1$

m : nombre quantique magnétique, m est un nombre d'entier qui détermine le nombre d'orbitales associées à un niveau d'énergie donné.

Le nombre d'orbitales associées à un niveau d'énergie donné est égal au nombre d'entiers compris entre $-l$ et l , y compris 0 .

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