

Unit 11: Chapter 7 - Chemical Reactions Review Answer Key

Define:

1. **Reactant:** *A substance that undergoes change in a chemical reaction*
2. **Product:** *New substances formed as a result of a chemical reaction.*
3. **Catalyst:** *A substance that affects the rate of a chemical reaction without being used up in the reaction.*
4. **Coefficient:** *Members that appear before a formula in a chemical equation to show the relative proportions of each reactant and product*
5. **Subscript:** *A number written lower in a chemical formula indicating the amount of atoms present of an element (this number is not to be changed once written in chemical formula)*
6. **Law of Conservation of Mass:** *Matter cannot be created nor destroyed; only changed (atoms can be rearranged)*
7. **Synthesis reaction (include an example):** *A chemical reaction in which two or more substances react to form a single substance.* $(A + B \rightarrow AB)$
8. **Decomposition reaction (include an example):** *A chemical reaction in which a compound breaks down into two or more simpler substances.* $(AB \rightarrow A + B)$
9. **Single replacement (include an example):** *A chemical reaction in which one element takes the place of another element in a compound.* $(AB + C \rightarrow AC + B)$
10. **Double replacement (include an example):** *A chemical reaction in which two compounds exchange positive ions and form two new compounds.* $(AB + CD \rightarrow AD + CB)$
11. **Combustion (include an example):** *A chemical reaction in which a substance reacts rapidly with oxygen, often producing heat and light.* $(C_2H_6 + O_2 \rightarrow CO_2 + H_2O)$
12. **Endothermic reaction:** *A chemical reaction that **absorbs** energy from its surroundings.*
13. **Exothermic reaction:** *A chemical reaction that **releases** energy to its surroundings.*
14. **Chemical equation (include an example):** *A representation of a chemical reaction in which the reactants and products are expressed as formulas.*
15. A **balanced equation** is an equation in which the number of atoms on the reactants' side **equals** the number of atoms on the products' side of the chemical equation.
16. **Equilibrium:** *A state in which the forward and reverse paths of a physical or chemical change take place at the same rate.*

Chapter 7 Chemical Reactions Answer Key

John T. Moore



Chapter 7 Chemical Reactions Answer Key:

Chemistry (Teacher Guide) Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books. The teacher is one who guides students through the subject matter, helps each student stay on schedule and be organized, and is their source of accountability along the way. With that in mind, this guide provides additional help through the laboratory exercises as well as lessons, quizzes, and examinations that are provided along with the answers. The lessons in this study emphasize working through procedures and problem solving by learning patterns. The vocabulary is kept at the essential level. Practice exercises are given with their answers so that the patterns can be used in problem solving. These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college. Guided labs are provided to enhance instruction of weekly lessons. There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions. It is important to see the hand of God and His principles and wisdom as it plays out in chemistry. This course integrates what God has told us in the context of this study. Features: Each suggested weekly schedule has five easy-to-manage lessons that combine reading and worksheets. Worksheets, quizzes, and tests are perforated and three-hole punched; materials are easy to tear out, hand out, grade, and store. Adjust the schedule and materials needed to best work within your educational program. Space is given for assignments, dates. There is flexibility in scheduling. Adapt the days to your school schedule. Workflow: Students will read the pages in their book and then complete each section of the teacher guide. They should be encouraged to complete as many of the activities and projects as possible as well. Tests are given at regular intervals with space to record each grade. About the Author: DR DENNIS ENGLIN earned his bachelor's from Westmont College, his master of science from California State University, and his EdD from the University of Southern California. He enjoys teaching animal biology, vertebrate biology, wildlife biology, organismic biology, and astronomy at The Master's University. His professional memberships include the Creation Research Society, the American Fisheries Association, Southern California Academy of Sciences, Yellowstone Association, and Au Sable Institute of Environmental Studies.

Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day. Relevant content, lively explorations, and a wealth of hands-on activities take students' understanding of science beyond the page and into the world around them. Now includes even more technology tools and activities to support differentiated instruction.

Physical Science C. Lon Enloe, Elizabeth Garnett, Jonathan Miles, Stephen Swanson, 2000-12-25 This is an introductory book that provides students with the tools to master the basic principles of physics and chemistry needed by the aspiring technology professional. Like all the books in the critically acclaimed Preserving the Legacy series, each chapter is divided into subsections featuring learning objectives and a

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General Chemistry, Reactions First Kevin Revell, 2024-12-04 Revell's General Chemistry empowers students to grasp essential topics and concepts with more ease Using a friendly approach the text uses metaphors and relatable examples to demystify even the most challenging subjects in general chemistry

General and Inorganic Chemistry Irena Kostova, 2025-04-21 General and Inorganic Chemistry covers the fundamental principles and general directions of chemistry and the physical and chemical properties of the elements and their compounds with an emphasis on their biological role The first part of the textbook presents basic theoretical topics such as the structure of the atom periodic table and law chemical bonding and complex compounds It includes topics related to chemical processes such as chemical thermodynamics chemical kinetics catalysis chemical equilibrium redox processes physicochemical analysis as well as topics on solutions such as disperse systems electrolyte solutions and colloidal solutions This part gives students systematic theoretical and practical knowledge in the field of general chemistry with an emphasis on biochemical processes The second part of the textbook is dedicated to chemical elements It is built on the concept of interconnection place in the periodic table chemical properties biological role of chemical elements and their compounds and is adapted to the needs of pharmaceutical practice It includes an analysis of the sources and preparations of the elements their common compounds their physical and chemical properties and their applications Attention is specifically focused on the role and influence of chemical elements and their compounds on biological systems and mainly on the human body Students are expected to build the necessary thinking and skills to apply this knowledge in their professional realization The compulsory course in general and inorganic chemistry is in line with the modern requirements for in depth fundamental knowledge and practical skills in the training of pharmacy and medical students At the same time the students pursuing MSc Chemical Engineering and other professional studies will also find the book extremely useful The objective is to provide the students with comprehensive treatment of the subject on modern lines

Gas Age-record, 1929 **Principles of Human Physiology** Ernest Henry Starling, 1926 **Addison-Wesley Chemistry**, 1990 *Gas Age*, 1929 Includes summaries of proceedings and addresses of annual meetings of various gas associations L C set includes an index to these proceedings 1884 1902 issued as a supplement to Progressive age Feb 15 1910 **Pharmaceutical Journal**, 1900 **General Chemistry** James E. Brady, Gerard E. Humiston, 1986 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and

streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems

Introductory Chemistry Steve Russo, Mike Silver, 2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support on The Chemistry Place™ Special Edition and on The Chemistry of Life CD ROM For college instructors and students

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