

## 4

Atomic Structure  
Reviewsheet

## A. Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this chapter. Each blank can be completed with a term, short phrase, or number.

Atoms of each element are 1 from the atoms of all other elements. Dalton theorized that atoms are indivisible, but the discovery of 2 particles changed this theory. We now know that atoms are made up of electrons, which have a 3 charge; 4, which have a positive charge; and 5, which are neutral. The latter two particles are found in the 6 of the atom.

It was 7 who discovered the nucleus of the atom. The nucleus has a 8 charge and it occupies a very small volume of the atom. In contrast, the negatively charged 9 occupy most of the volume of the atom.

The number of 10 in the nucleus of the atom is the atomic 11 of that element. Because atoms are electrically neutral, the number of protons and 12 in an atom are equal. The sum of the 13 and neutrons is the mass number. Atoms of the same element are identical in most respects, but they can differ in the number of 14 in the nucleus. Atoms that have the same number of protons but different mass numbers are called 15.

The 16 of an element is the weighted average of the masses of the isotopes of that element. Two isotopes of sulfur are  $^{32}_{16}\text{S}$  and  $^{34}_{16}\text{S}$ . An atom of the sulfur-32 isotope contains 17 protons and 18 neutrons. The sulfur-34 isotope has 19 protons and 20 neutrons.

1. different 41
2. sub atomic 42
3. negative 42
4. protons 42
5. neutrons 42
6. nucleus 43
7. Rutherford 43
8. positive 43
9. electrons 43
10. protons (elec) 44
11. number 44
12. electrons 44
13. protons (elec) 45
14. neutrons 45
15. isotopes 45
16. atomic mass 47
17. 16 45
18. 16 46
19. 16 46
20. 18 45

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**Thomas Evangelist**



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## **Table of Contents Chemistry Atomic Structure Chapter Review Answers**

1. Understanding the eBook Chemistry Atomic Structure Chapter Review Answers
  - The Rise of Digital Reading Chemistry Atomic Structure Chapter Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Atomic Structure Chapter Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Atomic Structure Chapter Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Atomic Structure Chapter Review Answers
  - Personalized Recommendations
  - Chemistry Atomic Structure Chapter Review Answers User Reviews and Ratings
  - Chemistry Atomic Structure Chapter Review Answers and Bestseller Lists
5. Accessing Chemistry Atomic Structure Chapter Review Answers Free and Paid eBooks
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  - Chemistry Atomic Structure Chapter Review Answers eBook Subscription Services

- Chemistry Atomic Structure Chapter Review Answers Budget-Friendly Options
- 6. Navigating Chemistry Atomic Structure Chapter Review Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Atomic Structure Chapter Review Answers Compatibility with Devices
  - Chemistry Atomic Structure Chapter Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Atomic Structure Chapter Review Answers
  - Highlighting and Note-Taking Chemistry Atomic Structure Chapter Review Answers
  - Interactive Elements Chemistry Atomic Structure Chapter Review Answers
- 8. Staying Engaged with Chemistry Atomic Structure Chapter Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Atomic Structure Chapter Review Answers
- 9. Balancing eBooks and Physical Books Chemistry Atomic Structure Chapter Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Atomic Structure Chapter Review Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Atomic Structure Chapter Review Answers
  - Setting Reading Goals Chemistry Atomic Structure Chapter Review Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Atomic Structure Chapter Review Answers
  - Fact-Checking eBook Content of Chemistry Atomic Structure Chapter Review Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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