

ESI Phone

User's Guide

FOR ESI Communications Servers

ESI-1000 • ESI-600 • ESI-200 • ESI-100 • ESI-50

Covering these current ESI phone models:

ESI 60 Business Phone • ESI 40 Business Phone • ESI 30D Business Phone
48-Key Feature Phone • ESI Cordless Handset II

. . . and these legacy ESI phone models:

24-Key Feature Phone • 12-Key Feature Phone • [Original] ESI Cordless Handset

0450-1047
Rev. R

For on-line help, visit www.esi-estech.com/support.

About ESI

ESI (Estech Systems, Inc.) designs and manufactures high-performance phone systems for businesses and organizations. ESI uses advanced technology to design IP and digital communications systems that integrate built-in capabilities, advanced features, and highly differentiated applications into flexible products that are easy to use and keep employees productive. ESI has sold over 250,000 business communications systems through hundreds of factory-trained Certified Resellers. Founded in 1987, ESI is a privately held corporation with headquarters in Plano, Texas.

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} + 2\text{H}^{+} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{H}_2\text{O}$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} + 2\text{CN}^{-} + 2\text{e}^{-}$ multiplied by a factor of 2 $2\text{e}^{-} + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^{-}$ 3 the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance

being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni²⁺ + 2e⁻ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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reactions $6e^{-} + 14H^{+} + Cr_2O_7^{2-} + 2Cr^{3+} + 7H_2O + Fe^{2+} + Fe^{3+} + e^{-} + 2$ equalize the electrons $6e^{-} + 14H^{+} + Cr_2O_7^{2-} + 2Cr^{3+} + 7H_2O + 6Fe^{2+} + 6Fe^{3+} + 6e^{-}$ multiplied by a factor of 6

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $C_6H_8O_6 + 2H_2NO_2 + C_6H_6O_6 + 2H_2O + 2NO$ choose 1 answer $C_6H_8O_6$ and NO_2 a $C_6H_8O_6$ and NO_2 b NO_2 and $C_6H_8O_6$ b NO_2 and $C_6H_8O_6$ h and NO_2 c h and