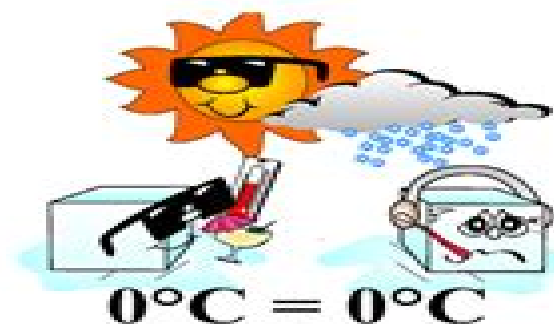


Heating/Cooling Curve

As a substance is heated, its particles begin to move faster and spread apart. The speed of the particles is related to their kinetic energy. The relative position of the particles is related to their potential energy. As solids, liquids, and gases are heated, most of the energy that is absorbed is converted to kinetic energy, and the temperature goes up. But as a substance melts or vaporizes, its particles spread out tremendously. As a result, the energy absorbed produces changes in the potential energy of the particles, so the temperature does not change as the phase changes. For that reason, the freezing point and the melting point of a substance are the same.



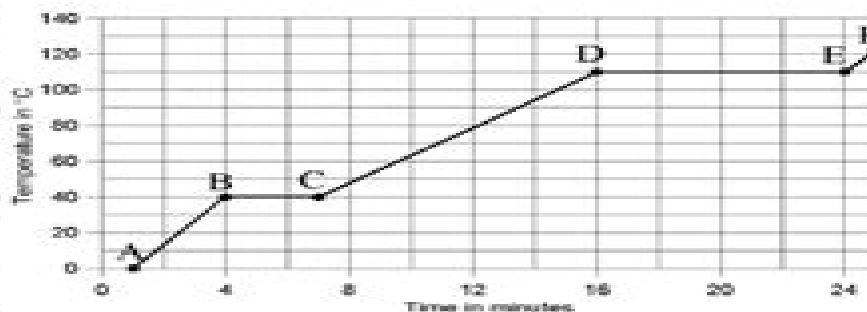
Base your answers to the following questions on the graph below which shows 10.0 kg of a substance that is solid at 0°C and is heated at a constant rate of 60 kilojoules per minute.

40°C
B-C
D-E

1. What is the temperature at which the substance can be both in the solid and the liquid phase?
2. During which lettered intervals is the internal potential energy of the substance increasing?

A-B
A-B
C-D
E-F

3. During which lettered intervals is the kinetic energy of the particles increasing?
4. How much heat is added to the substance from the time it stops melting to the time that it begins to boil?



540 kJ

5. What is the total heat needed to melt the substance (starting at time 0)?

1440 kJ

6. What is the total heat needed to vaporize the substance (starting at time 0)?

48 kJ/kg

7. What is the heat of vaporization of the substance?

A-B

8. During which lettered intervals is the substance solid?

C-D

9. During which lettered intervals is the substance in the liquid phase?

E-F

10. During which lettered intervals is the substance in the vapor phase?

110°C

11. What is the temperature at which the substance can be both in the liquid and the vapor phase?

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Table of Contents Evan P Silberstein 2003 Answer Key

1. Understanding the eBook Evan P Silberstein 2003 Answer Key
 - The Rise of Digital Reading Evan P Silberstein 2003 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Evan P Silberstein 2003 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Evan P Silberstein 2003 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Evan P Silberstein 2003 Answer Key
 - Personalized Recommendations
 - Evan P Silberstein 2003 Answer Key User Reviews and Ratings
 - Evan P Silberstein 2003 Answer Key and Bestseller Lists

5. Accessing Evan P Silberstein 2003 Answer Key Free and Paid eBooks
 - Evan P Silberstein 2003 Answer Key Public Domain eBooks
 - Evan P Silberstein 2003 Answer Key eBook Subscription Services
 - Evan P Silberstein 2003 Answer Key Budget-Friendly Options
6. Navigating Evan P Silberstein 2003 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Evan P Silberstein 2003 Answer Key Compatibility with Devices
 - Evan P Silberstein 2003 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Evan P Silberstein 2003 Answer Key
 - Highlighting and Note-Taking Evan P Silberstein 2003 Answer Key
 - Interactive Elements Evan P Silberstein 2003 Answer Key
8. Staying Engaged with Evan P Silberstein 2003 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Evan P Silberstein 2003 Answer Key
9. Balancing eBooks and Physical Books Evan P Silberstein 2003 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Evan P Silberstein 2003 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Evan P Silberstein 2003 Answer Key
 - Setting Reading Goals Evan P Silberstein 2003 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Evan P Silberstein 2003 Answer Key
 - Fact-Checking eBook Content of Evan P Silberstein 2003 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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