

Vol. of 1 mol of ideal gas at STP/NTP =  $22.4 \text{ L}$

उत्स नान एरिया  
- 3 same sabse kam significant figure hoga pt. ke bad utne hi hoga ans me.

**\* Significant fig**  
 0 to the left of the 1st non 0 digit are sig. not sig.  
 eg. 0.0012  $\Rightarrow$  2  
 Terminal 0 are not sig. if there is no decimal pt.  
 eg. 1000  $\Rightarrow$  1 fig.  
 0 at end is sig. if there is decimal pt.  
 Exponential term are not considered sig.

अंश - 3/131  
 $a \times b = c$  } - add/subtract is least amount  
 $\Rightarrow$  3 same sabse kam fig hoga vo sig. fig ans me hoga

**\* LAW OF MASS CONSV**  
 In a react<sup>n</sup>  $\rightarrow$  actual mass of reactant = mass of prod.  
 Rasm me zitna use hua.

Except<sup>n</sup>  $\rightarrow$  Nuclear Rxn (C = mc<sup>2</sup>)  
 PppWed<sup>n</sup>  $\rightarrow$  Balanced Rxn

# MOLE CONCEPT

**\* LAW OF CONST PROPORTN**  
 Elements in a comp are in fixed ratio by mass.  
 eg.  $\text{CO}_2$   
 $\begin{matrix} 12 \\ \diagdown \\ \text{C} \\ \diagup \\ 32 \\ \text{O}_2 \end{matrix} \Rightarrow 3:8$   
 Not valid for isomer.  
 $\% \text{ of C} = \frac{3}{11} \times 100$

Q. cal. the mass of C req to form 132g of  $\text{CO}_2$ .  
 $\Rightarrow \text{C} + 2\text{O} \rightarrow \text{CO}_2$   
 $\begin{matrix} 12 & & 32 \\ \text{C} & & \text{O}_2 \end{matrix} \Rightarrow \begin{matrix} 44 \\ \text{CO}_2 \end{matrix}$   
 44g  $\text{CO}_2$  ke lie 12g C  
 $\therefore 132\text{g}$  ke lie  
 $= \frac{12}{44} \times 132 = 36\text{g}$

**\* Law of Multiple Proport<sup>n</sup>**  
 elements should be same or same fixed mass  
 eg.  $\text{CO}$  &  $\text{CO}_2$  1:2 ratio  
 Other eg.  $\text{N}_2\text{O}$ ,  $\text{N}_2\text{O}_3$ ,  $\text{N}_2\text{O}_5$   
 except<sup>n</sup>  $\rightarrow$   $\text{N}_2\text{O}_3$  &  $\text{N}_2\text{O}_5$   
 Avgad<sup>n</sup> vo law

**\* Same (P, T, R, V) of molecules of gases** some (mass eg. 1 mol of)  
 $\text{X}_2$ ,  $\text{X}_3$ ,  $\text{X}_4$   
 $2 : 3 : 4$  molecules  
 $2 : 3 : 4$  atoms  
 [different]

**\* Gay Lussac's law**  
 CONDINS: +T, P some  
 at least 2 gases.  
 $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$   
 $1 : 1 : 2$   
 Vol. K. Lie

**\* DALTON'S ATOMIC THEORY**  
 Atoms  $\rightarrow$  Neither created nor destroyed

**\* MOLE CONCEPT**  
 $\frac{A}{Z} \times (P+N)$   
 Atomic No. (P)  
 Mass of Proton =  $1.67 \times 10^{-27} \text{ kg}$   
 Neutron =  $1.67 \times 10^{-27} \text{ kg}$   
 Electron =  $9.1 \times 10^{-31} \text{ kg}$   
 MP  $\approx 1837$   $1/e$

**\* Atomic mass unit (amu)**  
 $1 \text{ amu} = \frac{1}{12} \times \text{Mass of } ^{12}\text{C}$   
 $1 \text{ amu} = 1.67 \times 10^{-27} \text{ kg}$   
 eg. Mass of  $^{12}\text{C}$  = 12 amu  
**\* Relative Atomic mass**  
 Unitless  
 eg. mass. A.M. of C = 12

**\* Molecular Mass**  
 $\text{H}_2\text{SO}_4 = 98 \text{ amu}$   
 $\rightarrow$  Mass of 1 molecule of  $\text{H}_2\text{SO}_4$

**\* FORMULA MASS**: for ionic comp.  
**\* MOLE**  
 Amount of substance.  
 1 mole = No particles  $(6.022 \times 10^{23})$   
 $1 \text{ amu} \leq 1.67 \times 10^{-27} \text{ kg}$   
 $\text{Na} = 6.022 \times 10^{23}$   
 $1 \text{ amu} \times \text{Na} = 1 \text{ g}$

# Mole Concept Theory Notes Pdfslibforme

**Yicheng Fang**



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