

KEY

3. The pH in a 0.25 M solution of the acid HBrO is 4.65. Using this, calculate the value of Ka for the acid HBrO. (4 marks) $pH = 4.65$ so $[H_3O^+] = \text{antilog}(-4.65) = 2.239 \times 10^{-5} M$



[I]	0.25		0	0
[C]	-2.239×10^{-5}		$+2.239 \times 10^{-5}$	$+2.239 \times 10^{-5}$
[E]	$0.25 - 2.239 \times 10^{-5}$		2.239×10^{-5}	2.239×10^{-5}

4

$$K_a = \frac{[H_3O^+][BrO^-]}{[HBrO]}$$

$$= \frac{(2.239 \times 10^{-5})^2}{(0.25 - 2.239 \times 10^{-5})} = 2.0 \times 10^{-9}$$

Answer $K_a = 2.0 \times 10^{-9}$

4. The pH in a solution of benzoic acid is 2.355. Determine the molar concentration of the benzoic acid. (4 marks) $pH = 2.355$ $[H_3O^+] = \text{antilog}(-2.355) = 0.004416 M$



[I]	C_0		0	0
[C]	-0.004416		$+0.004416$	$+0.004416$
[E]	$C_0 - 0.004416$		0.004416	0.004416

4

$$K_a = \frac{[H_3O^+][C_6H_5COO^-]}{[C_6H_5COOH]}$$

$$K_a = \frac{(0.004416)^2}{(C_0 - 0.004416)}$$

$$C_0 - 0.004416 = \frac{(0.004416)^2}{6.5 \times 10^{-5}} = 0.299976$$

$$C_0 = 0.299976 + 0.004416 = 0.30439 M$$

Answer $C_0 = 0.30 M$

5. Find the value of Kb for the oxalate ion ($C_2O_4^{2-}$). (1 mark)

$$K_b(C_2O_4^{2-}) = \frac{K_w}{K_a(HC_2O_4^-)} = \frac{1.00 \times 10^{-14}}{6.4 \times 10^{-5}} = 1.56 \times 10^{-10}$$

1

Answer $K_b = 1.6 \times 10^{-10}$

6. The value of Kb for the weak base methylamine (CH_3NH_2) is 4.4×10^{-4} . Calculate the value of Ka for the acid $CH_3NH_3^+$. (1 mark)

$$K_a(CH_3NH_3^+) = \frac{K_w}{K_b(CH_3NH_2)} = \frac{1.00 \times 10^{-14}}{4.4 \times 10^{-4}} = 2.27 \times 10^{-11}$$

1

Answer $K_a = 2.3 \times 10^{-11}$

10

Basic Stoichiometry Phet Lab Answer Key

D Kirk



Basic Stoichiometry Phet Lab Answer Key :

The book delves into Basic Stoichiometry Phet Lab Answer Key . Basic Stoichiometry Phet Lab Answer Key is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Basic Stoichiometry Phet Lab Answer Key , encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Basic Stoichiometry Phet Lab Answer Key
- Chapter 2: Essential Elements of Basic Stoichiometry Phet Lab Answer Key
- Chapter 3: Basic Stoichiometry Phet Lab Answer Key in Everyday Life
- Chapter 4: Basic Stoichiometry Phet Lab Answer Key in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Basic Stoichiometry Phet Lab Answer Key . The first chapter will explore what Basic Stoichiometry Phet Lab Answer Key is, why Basic Stoichiometry Phet Lab Answer Key is vital, and how to effectively learn about Basic Stoichiometry Phet Lab Answer Key .

3. In chapter 2, this book will delve into the foundational concepts of Basic Stoichiometry Phet Lab Answer Key . This chapter will elucidate the essential principles that need to be understood to grasp Basic Stoichiometry Phet Lab Answer Key in its entirety.

4. In chapter 3, the author will examine the practical applications of Basic Stoichiometry Phet Lab Answer Key in daily life. This chapter will showcase real-world examples of how Basic Stoichiometry Phet Lab Answer Key can be effectively utilized in everyday scenarios.

5. In chapter 4, this book will scrutinize the relevance of Basic Stoichiometry Phet Lab Answer Key in specific contexts. This chapter will explore how Basic Stoichiometry Phet Lab Answer Key is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Basic Stoichiometry Phet Lab Answer Key . The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Basic Stoichiometry Phet Lab Answer Key .

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