

Chapter 9: Chemical Reactions

- _____ 1. Which of the following could be an indication of a chemical reaction of a compound or element?
- a. It is a solid at room temperature.
 - b. It melts at 631°C .
 - c. It expands upon freezing.
 - d. It burns in an atmosphere of chlorine.
- _____ 2. In this unbalanced chemical reaction, which components are the reactants?
- $$\text{C}_7\text{H}_{16} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$$
- a. C_7H_{16} and O_2
 - b. C_7H_{16} and CO_2
 - c. C_7H_{16} and H_2O
 - d. CO_2 and H_2O
- _____ 3. When $\text{Ca}(\text{OH})_2$ reacts with H_2SO_4 , what are the products.
- a. $\text{Ca}(\text{OH})_2$ and H_2SO_4
 - b. CaH_2 and H_2SO_4
 - c. CaSO_4 and H_2O
 - d. CaO and H_2SO_4
- Hint: This is a double-replacement reaction, positive ions switch places.
- _____ 4. What are the products of the reaction between barium hydroxide and hydrochloric acid?
- a. Barium chloride and water
 - b. Barium hydroxide and water
 - c. Barium chloride and barium hydride
 - d. Chloric acid and barium hydroxide
- Hint: Write out the skeleton equation. This is a double-replacement reaction.
- _____ 5. This symbol separates the reactants from products and means "to yield".
- a. +
 - b. (s)
 - c. \longrightarrow
 - d. (g)
- _____ 6. In a chemical equation, the number written in front of the chemical formula is the
- a. Reactant
 - b. Subscript
 - c. Equation
 - d. Coefficient
- _____ 7. When there are exactly the same number and type of atoms on both sides of a chemical equation, the equation is _____.
- a. Balanced
 - b. Ionic
 - c. Synthetic
 - d. Decomposed
- _____ 8. The chemical equation $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$ is classified as a
- a. Synthesis reaction
 - b. Decomposition reaction
 - c. Single-replacement reaction
 - d. Double-replacement reaction
- _____ 9. In the combustion reaction of propane (C_3H_8) with oxygen gas, what are the products?
- a. C_3H_8 and O_2
 - b. C_3H_8 and CO_2
 - c. O_2 and H_2O
 - d. H_2O and CO_2
- _____ 10. The major physical indicator(s) of a combustion reaction is

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Ingo Waldemar Dagobert Hackh



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Chemical Reactions Denise Walker, 2007 This title introduces the reader to the huge variety of chemical reactions that shape our world Find out all about explosions learn about how to start reactions and understand how chemical equations work

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