

Section 12.1

OBJECTIVE 1

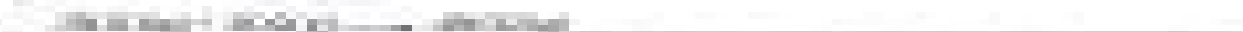
Write complete balanced redox reactions that are indicated.

1. $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$ (acidic solution)
 - (a) identify the oxidizing and reducing agents
 - (b) balance the reaction in acidic solution
 - (c) identify the oxidized and reduced species
2. $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$ (basic solution)
 - (a) identify the oxidizing and reducing agents
 - (b) balance the reaction in basic solution
 - (c) identify the oxidized and reduced species
3. $\text{Fe}^{2+} + \text{H}_2\text{O}_2 \rightarrow \text{Fe}^{3+} + \text{H}_2\text{O}$ (neutral solution)
 - (a) identify the oxidizing and reducing agents
 - (b) balance the reaction in neutral solution
 - (c) identify the oxidized and reduced species

4. Write the net ionic equation for the following redox reaction:



a. Balance the equation:



b. Write the net ionic equation for the reaction.

c. In which of the following solutions would the reaction be most likely to occur?

d. Write a balanced equation for the reaction in the solution you chose.

e. Write a balanced equation for the reaction in the solution you chose.

Chapter 9 Stoichiometry Section 1 Answers

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