

## Periodic Table Trend Activities:



**Purpose:** The Periodic Table is a useful tool that helps us predict properties of various elements. In this activity, we will look at 3 properties:

1. **Atomic radius:** The distance from the atomic nucleus to the outmost occupied orbital in a atom
2. **Electronegativity:** The ability of an atom to attract a bonded pair of electrons from another atom
3. **Ionization Energy:** The energy needed to remove an electron from a neutral gaseous atom
4. **Electron Affinity:** The energy released when an electron is added to a neutral gaseous atom

**Procedure:** Click on the [Cool Periodic Table link](#) (click on the element symbol). Record the atomic radius, electronegativity, and first ionization energy for each element in the table below:

<http://www.lyncsburg.net/hhs/chemistry/trends/>

**Data:**

Element	Atomic Radius	Electronegativity	1 <sup>st</sup> Ionization Energy
Li			
Be			
B			
Na			
Mg			
Al			
C			
Si			
N			
P			

**Discussion:**

1. What appears to be the trend in atomic radius as you move from left to right in a row?
2. What appears to be the trend in atomic radius as you move down a column?

# Chemistry Lab Periodic Trends Computer Activity Answers

**L Cohen**



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*A Study of Silicon-29 Chemical Shift Anisotropy in Disilenes and Cyclic Organosilicon Compounds ; And, General Chemistry Instruction by Computer* Jefferson D. Cavalieri,1993     **Prentice Hall Science Explorer** Michael J. Padilla,2002  
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