

9/2/16

S. Sai Krishna

Date \_\_\_\_\_  
Page \_\_\_\_\_  
ClassmateThermochemistry

$$Q = \Delta E - W$$

$$Q = \Delta E - (-P\Delta V)$$

$$Q = \Delta E + P\Delta V$$

- At constant Volume.

$$Q = \Delta E$$

- At const. Pressure.

$$Q_p = \Delta E + P\Delta V$$

$$H = \Delta E + P\Delta V$$

$$H = \Delta E + \Delta nRT$$

$\Delta n =$  no. of moles of Prod  
- no. of moles of react.

→ For  $\Delta H$ .

- If  $\Delta H$  is -ve. reaction is exothermic.
- If  $\Delta H$  is +ve. reaction is endothermic.

Q. The heat of combustion of ethane at 17°C and const volume is  $-322.1 \text{ kJ}$ . Calculate the heat of combustion at constant pressure.



# Engineering Chemistry Full Notes

**A Gutmann**



## **Engineering Chemistry Full Notes:**

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In a fast-paced digital era where connections and knowledge intertwine, the enigmatic realm of language reveals its inherent magic. Its capacity to stir emotions, ignite contemplation, and catalyze profound transformations is nothing lacking extraordinary. Within the captivating pages of **Engineering Chemistry Full Notes** a literary masterpiece penned by a renowned author, readers set about a transformative journey, unlocking the secrets and untapped potential embedded within each word. In this evaluation, we shall explore the book's core themes, assess its distinct writing style, and delve into its lasting affect the hearts and minds of those who partake in its reading experience.

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