

## Lab 22A

### Investigating Chemical Equilibrium

**Purpose:** To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a function of concentration and temperature and relate the to systems, LeChâtelier's principle, and equilibrium constants applying LeChâtelier's principle.

#### Proble:

1. Write a balanced equation for the reaction in part 1.  
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
2. Write a balanced equation for the reaction in part 2.  
 $\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$
3. Write a balanced equation for the reaction in part 3.  
 $\text{Co}(\text{H}_2\text{O})_6^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$
4. In part 1, what is the conjugate acid and base? Write the  $K_{\text{eq}}$  for the reaction. Is the reaction endothermic or exothermic? Explain the color change. How does the color change as a function of temperature? Write the  $K_{\text{eq}}$  for the reaction. Is the reaction endothermic or exothermic? Explain the color change.
5. Write the  $K_{\text{eq}}$  for the reaction in part 2.  
 $K_{\text{eq}} = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^{-}]}$
6. Write the  $K_{\text{eq}}$  for the reaction in part 3.  
 $K_{\text{eq}} = \frac{[\text{CoCl}_4^{2-}]}{[\text{Co}(\text{H}_2\text{O})_6^{2+}][\text{Cl}^{-}]^4}$
7. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$
8. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$
9. Write the equation for the reaction between  $\text{Fe}^{3+}$  and  $\text{OH}^{-}$  ions.  
 $\text{Fe}^{3+} + 3\text{OH}^{-} \rightleftharpoons \text{Fe}(\text{OH})_3(\text{s})$

# Investigating Chemical Equilibrium Lab 12a Answer

**Society of Chemical Industry (Great  
Britain)**



## Investigating Chemical Equilibrium Lab 12a Answer:

*Theoretical Principles of Isotope Methods for Investigating Chemical Reactions* Simon Zalmanovich Roginskiĭ,1956

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**Aerospace Reports** ,1970 **General Chemistry Laboratory Text with Qualitative Analysis** Wilbert Hutton,1968

*Journal of the Society of Chemical Industry* Society of Chemical Industry (Great Britain),1907 **Journal of the Society of Chemical Industry** ,1907 **Quantitative Chemical Analysis** Daniel C. Harris,2010-04-30 QCA is the bestselling textbook of choice for analytical chemistry It offers a modern portrait of the techniques of chemical analysis backed by a wealth of real world applications This edition features new coverage of spectroscopy and statistics new pedagogy and enhanced lecturer support Technical Abstract Bulletin ,1965 **Physics Briefs** ,1993 **Chemical Abstracts** ,1908

**Chemical, Color and Oil Record** ,1925 **Journal of the American Chemical Society** American Chemical Society,1925 Proceedings of the Society are included in v 1 59 1879 1937 **An Introductory Guide to EC Competition Law and Practice** Valentine Korah,1994 **Energy Research Abstracts** ,1977 Semiannual with semiannual and annual indexes References to all scientific and technical literature coming from DOE its laboratories energy centers and contractors Includes all works deriving from DOE other related government sponsored information and foreign nonnuclear information Arranged under 39 categories e g Biomedical sciences basic studies Biomedical sciences applied studies Health and safety and Fusion energy Entry gives bibliographical information and abstract Corporate author subject report number indexes

**El-Hi Textbooks in Print** ,1978 *Technical Note* ,1912 **Journal** Society of Chemical Industry (Great Britain),1907

**New Technical Books** ,1926 *The Engineer* ,1882 **Bulletin of the Chemical Society of Japan** Nihon Kagakkai,2006

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