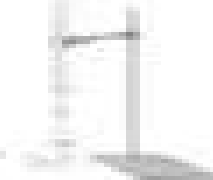




# AP Chem Investigation 12 – Redox Titration Lab



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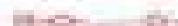
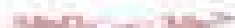
- Check to be sure you have at least 20 mL of  $K_2Cr_2O_7$  in your buret. If not, add some  $K_2Cr_2O_7$  to your buret.
- Be sure there are not air bubbles in the buret. If there are, stretch some  $K_2Cr_2O_7$  out of the buret into a waste beaker until the air bubble comes out. You can dump this solution down the sink with running water.
- Take an initial reading of  $K_2Cr_2O_7$  from the buret. Record this value in the data table.
- Measure out exactly 1.00 mL of  $H_2O_2$  using a graduated pipet and add it to a clean flask.
- Add 20 mL of 1M  $H_2SO_4$  to the flask using a graduated cylinder. This is so that the titration can occur in acidic conditions.
- Add water to the flask to bring the solution up to about 50-60 mL of volume. (This is just to improve your ability to see the color change at the endpoint. Adding water will NOT affect the number of moles of  $H_2O_2$  present in the flask, so it won't affect the volume of titrant required to reach the equivalence point.)
- Titrate the  $H_2O_2$  to the endpoint (until a first good color persists). Try not to add even just a drop of  $K_2Cr_2O_7$  past the endpoint! Record the volume of titrant.
- Dump your solution down the sink. Leave the leftover  $K_2Cr_2O_7$  in your buret for the next class.
- Check the label on the  $K_2Cr_2O_7$  bottle and record the molarity of the  $K_2Cr_2O_7$  solution in your data section.

### Data Table

Volume of $H_2O_2$ sampled	1.00 mL
Initial reading of $K_2Cr_2O_7$ volume on buret	19.11 mL
Final reading of $K_2Cr_2O_7$ volume on buret	17.41 mL
Concentration of $K_2Cr_2O_7$ solution on bottle	0.00880 M

### Questions

- Balance the equation below using the half-reaction method. Show all work.



- What volume of  $K_2Cr_2O_7$  was added to reach the endpoint?

$$17.41 - 19.11 = 1.70 \text{ mL}$$

- Knowing the volume and concentration of  $K_2Cr_2O_7$ , how many moles of  $K_2Cr_2O_7$  were added to reach the endpoint?

$$1.70 \text{ mL} \times (0.00880 \text{ mol/L}) \times (1000 \text{ mL/L}) = 0.01496 \text{ mol}$$

- Using the balanced equation, how many moles of  $H_2O_2$  were in your sample? (Remember  $3K_2Cr_2O_7$  to  $3H_2O_2$  in the balanced equation.)

$$1.70 \times 10^{-2} \text{ mol } K_2Cr_2O_7 \times (3 \text{ mol } H_2O_2 / 3 \text{ mol } K_2Cr_2O_7) = 0.01496 \text{ mol } H_2O_2$$

- Convert the moles of  $H_2O_2$  into grams.

$$0.01496 \text{ mol } H_2O_2 \times (34.01 \text{ g/mol}) = 0.509 \text{ g}$$

- Imagine that your sample of  $H_2O_2$  was mostly water and contained very little  $H_2O_2$  molecules. If it was mostly water, it would be expected to have a density similar to water. The density of water is 1.00 g/mL. Let's assume the sample had a density of 1.00 g/mL. Using the mass of  $H_2O_2$  calculated in the previous question along with the fact that your sample was exactly 20.00 mL, calculate the percentage of your sample that was  $H_2O_2$ .

$$(0.509 / 20.00) \times 100\% = 2.54\%$$

- The label on the bottle states that the bottle of  $H_2O_2$  is a 3% solution. Compare this value to your calculation in the previous question. Is the label correct? Why or why not? (The 3% is based on  $H_2O_2$  breaking down over time.)

- If necessary to know the exact volume of...

a)  $H_2O_2$  sample added to the flask? 1.70 mL

b) water added to the flask? 38.30 mL

c)  $K_2Cr_2O_7$  added to the flask? 1.70 mL

# Ap Chem Titration Lab Answers

**JL Elias**



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**Illustrated Guide to Home Chemistry Experiments** Robert Bruce Thompson,2012-02-17 For students DIY hobbyists and science buffs who can no longer get real chemistry sets this one of a kind guide explains how to set up and use a home chemistry lab with step by step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work Purify alcohol by distillation Produce hydrogen and oxygen gas by electrolysis Smelt metallic copper from copper ore you make yourself Analyze the makeup of seawater bone and other common substances Synthesize oil of wintergreen from aspirin and rayon fiber from paper Perform forensics tests for fingerprints blood drugs and poisons and much more From the 1930s through the 1970s chemistry sets were among the most popular Christmas gifts selling in the millions But two decades ago real chemistry sets began to disappear as manufacturers and retailers became concerned about liability em The Illustrated Guide to Home Chemistry Experiments steps up to the plate with lessons on how to equip your home chemistry lab master laboratory skills and work safely in your lab The bulk of this book consists of 17 hands on chapters that include multiple laboratory sessions on the following topics Separating Mixtures Solubility and Solutions Colligative Properties of Solutions Introduction to Chemical Reactions Stoichiometry Reduction Oxidation Redox Reactions Acid Base Chemistry Chemical Kinetics Chemical Equilibrium and Le Chatelier s Principle Gas Chemistry Thermochemistry and Calorimetry Electrochemistry Photochemistry Colloids and Suspensions Qualitative Analysis Quantitative Analysis Synthesis of Useful Compounds Forensic Chemistry With plenty of full color illustrations and photos Illustrated Guide to Home Chemistry Experiments offers introductory level sessions suitable for a middle school or first year high school chemistry laboratory course and more advanced sessions suitable for students who intend to take the College Board Advanced Placement AP Chemistry exam A student who completes all of the laboratories in this book will have done the equivalent of two full years of high school chemistry lab work or a first year college general chemistry laboratory course This hands on introduction to real chemistry using real equipment real chemicals and real quantitative experiments is ideal for the many thousands of young people and adults who want to experience the magic of chemistry

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