

AP Chemistry: Bonding

Multiple Choice

41. Which of the following molecules has the shortest bond length?
 (A) N_2 (B) O_2 (C) Cl_2 (D) Br_2 (E) I_2

51. Pi bonding occurs in each of the following species EXCEPT...
 (A) CO_2 (B) C_2H_4 (C) CN^- (D) C_6H_6 (E) CH_4

60. Which of the following has a zero dipole moment?
 (A) HCN (B) NH_3 (C) SO_2 (D) NO_2 (E) PF_5

8. Use the following answers for questions 3 - 9.

- (A) A network solid with covalent bonding
- (B) A molecular solid with zero dipole moment
- (C) A molecular solid with hydrogen bonding
- (D) An ionic solid
- (E) A metallic solid

3. Solid ethyl alcohol, C_2H_5OH

9. Silicon dioxide, SiO_2

30. For which of the following molecules are resonance structures necessary to describe the bonding satisfactorily?

- (A) H_2S (B) SO_2 (C) CO_2 (D) OF_2 (E) PF_3

Hydrogen Halide	Normal Boiling Point, $^{\circ}C$
HF	19
HCl	-85
HBr	-67
HI	-35

18. The liquefied hydrogen halides have the normal boiling points given above. The relatively high boiling point of HF can be correctly explained by which of the following?

- (A) HF gas is more ideal
- (B) HF is the strongest acid.
- (C) HF molecules have a smaller dipole moment.
- (D) HF is much less soluble in water.
- (E) HF molecules tend to form hydrogen bonds.

42. The $SbCl_5$ molecule has trigonal bipyramid structure. Therefore, the hybridization of Sb orbitals should be...

- (A) sp^2 (B) sp^3 (C) sp^3d (D) sp^3d^2 (E) sp^3d^2

11. Use these answers for questions 11 - 14.

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RS Peters



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