

Lab 22A

Investigating Chemical Equilibrium

Purpose: To investigate the macroscopic properties of two chemical systems at equilibrium, observe shifts in equilibrium concentrations as a result of concentration and temperature and make use of Le Chatelier's principle and an engineering design process through applying Le Chatelier's principle.

Proble:

1. Write a balanced equation for the reaction in part 1.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
2. Write a balanced equation for the reaction in part 2.
$$\text{Fe}^{3+}(\text{aq}) + \text{SCN}^{-}(\text{aq}) \rightleftharpoons \text{FeSCN}^{2+}(\text{aq})$$
3. Write a balanced equation for the reaction in part 3.
$$\text{Co}(\text{H}_2\text{O})_6^{2+}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \text{CoCl}_4^{2-}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$$
4. In part 1, what is the conjugate acid and base? In part 2, what is the conjugate acid and base?
Conjugate acids have an extra proton and base is the other side of the reaction. The conjugate acid of the base is the species that formed in the reaction and the other side of the reaction is the conjugate base.
5. What is the conjugate acid and base of the reaction in part 3?
Water is added to the other side of the reaction to make water and hydroxide ions.
6. What is the conjugate acid and base of the reaction in part 4?
Hydroxide ion is the conjugate base and water is the conjugate acid.
7. Write the equation for the reaction between H^{+} and OH^{-} ions.
$$\text{H}^{+} + \text{OH}^{-} \rightleftharpoons \text{H}_2\text{O}$$
8. Write your own reaction that H^{+} and OH^{-} ions.
Sulfuric acid H_2SO_4 will react with hydroxide ions to produce bisulfate ions HSO_4^{-} and water H_2O as the products.
9. What is a spectator ion?
A spectator ion means in the reaction that the ion does not participate in the reaction. It is unreactive and does not affect equilibrium.

Investigating Chemical Equilibrium Lab 12a Answer

Marco Cascella



Investigating Chemical Equilibrium Lab 12a Answer:

Theoretical Principles of Isotope Methods for Investigating Chemical Reactions Simon Zalmanovich Roginskiĭ,1956

CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE Kevin Smith,2024-02-01 A comprehensive summary of Grade 11 12 Physics Simple logical summaries with example exam questions and work through solutions The book covers the fundamentals of Grade 11 12 Physics and complements the material in any class text **Scientific and Technical**

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