

Structure of H<sub>2</sub>O  
hydrogen bonding  
Denser at 4°C



# unit one guide

Biological molecules  
proteins  
nucleic acids

H<sub>2</sub>O provides cushion, adhesion, surface tension, high specific heat, high heat of vaporization, excellent solvent, stable chemistry, low compressibility

THIS GENE: when ions → water expands due to hydrogen bonding, molecules next to each other

hydrogen bonds require lots of energy to change state of matter

CONDENSATION → when two molecules → products hydrogen and hydroxyl ions, releases pH → more hydrogen

OH<sup>-</sup> → attracted to same  
OH<sup>-</sup> → attracted to other  
OH<sup>-</sup> → hydrogen bonds  
OH<sup>-</sup> → hydrogen bonds  
OH<sup>-</sup> → surface of the liquid  
OH<sup>-</sup> → surface of the liquid

<b>C</b>	<b>H</b>	<b>N</b>	<b>O</b>	<b>P</b>	<b>S</b>
Carbon	Hydrogen	Nitrogen	Oxygen	Phosphorus	Sulfur

more hydrogen  
C → build biological molecules, cell formation  
N → build proteins, nucleic acids  
P → build nucleic acids + lipids  
→ make up 98% of organisms

Carbohydrates ↓ C, H, O	Lipids ↓ C, H, O, P	Nucleic acids ↓ C, H, N, O, P	Proteins ↓ C, H, N, O, S
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Condensation synthesis → the removal of an H<sub>2</sub>O molecule to bond monomers

## Nucleic acid structure:

• the carbon sugar (deoxyribose/ribose), phosphate, nitrogenous base (A, T, C, G), exact shape of strand



Starch: 1,4 glucose  
Glycogen: 1,4 glucose  
Cellulose: 1,4 glucose  
Chitin: 1,4 glucose

Proteins: specific order of amino acids  
Primary structure

1° group: interactions determine structure + function

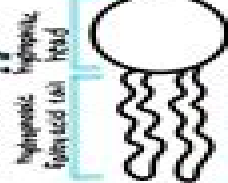


## PROTEINS

• amino acids bond via a peptide bond (covalent) through dehydration synthesis  
• 1° group change can cause a hydrophobic protein to lose the place of a hydrophilic protein

## membranes

• made up of phospholipids  
• create selectively permeable barrier  
• contain phospholipids, cholesterol, and proteins



## LIPIDS

Phospholipid macromolecules

• differences in saturation determine structure and function  
• Cells, sterols, phospholipids  
• Saturated → no double bonds  
• Unsaturated → has double bonds  
• Lipids: triglycerides, cholesterol, most of the membrane



Amino Acids: have directionality

• linear chains, connected by covalent bonds  
• amino end and carboxyl end

## PROTEIN STRUCTURE

• Primary: sequence of amino acids, specific order determine protein shape  
• Secondary: how folding of chain into alpha helix and beta sheet occurs  
• Tertiary: 3-D shape, hydrogen bonds, disulfide link, ionic bonds, hydrophobic interactions  
• Quaternary: interactions between multiple polypeptide chains

## + FUNCTION

• Enzymes → speeds up reaction rates in the cell  
• Cell structure → part of the cellular membrane  
• Communication between cells

## DENATURATION:

• the protein starts to lose its shape/structure when (bonds break)  
• Temperature change  
• pH of surrounding area  
• Salt concentration

NH<sub>2</sub> = Amino  
COOH = Carboxyl

## NUCLEIC ACIDS

• add new nucleotides to the 3' end  
• two nitrogenous bases: purine + pyrimidine  
• pyrimidine is single ringed, purine is double  
• Antiparallel double helix  
• 5' on phosphate end, 3' on hydroxyl end  
• both back and forth have nitrogenous bases (purines) + sugar-phosphate backbone

## DEHYDRATION SYNTHESIS



## ANTIPARALLEL



## Chaperones:

• protein molecules that assist protein folding

## ALDOSE VS KETOSE

• contains aldehyde functional group

## ALPHA VS BETA

• contains ketone functional group



## ALPHA VS BETA

• all monomers are in the same orientation



• Every glucose monomer is upside down from its neighbor

# Ap Biology Chapter 12 Notes

**R Pring**



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**Biomacromolecules** Mohamed Elzagheid,2024-12-30 The second edition of the Macromolecular Chemistry broadens into two areas biomacromolecules Volume 1 and polymers Volume 2 Biomacromolecules covers carbohydrates lipids proteins nucleic acids their classifications and properties The first chapter looks at the structural formulas and cyclic forms of monosaccharides as well as their synthesis and breakdown Cyclization enolization isomerization tautomerization mutarotation and epimerization are all briefly described Examples of disaccharides and polysaccharides are also presented The second chapter covers triglycerides steroids vitamins and their constituents The third chapter examines the primary structure of proteins including amino acid properties peptide bond formation and peptide synthesis It also addresses secondary tertiary and quaternary structures The book concludes with a chapter on nucleic acids which covers the chemistry of nucleosides and oligonucleotides as well as topics such as genetic code DNA secret code Polymerase Chain Reaction and DNA fingerprinting

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