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# ACS Gen Chem 1 Exam

**Ion** - A species of an element in which the number of electrons does not equal the number of protons.

**Isotope** - Each isotope of an atom has a different number of neutrons.

**Isoelectric** - Two atoms with the same charge.

**Isotopic** - Two atoms with the same number of neutrons.

**Transition Metal Orbitals** - The 4s orbital is filled before the 3d orbital, but listed after.

**Paramagnetic** - Atoms that possess a permanent magnetic charge due to the presence of unpaired electrons.

**Electropositive** - Tending to lose electrons and form positive ions.

**Tetrahedral Bond Angle** -  $109.5^\circ$

**Resonance Structures** - Two equally correct arrangements of one molecule's electrons; generally the molecule exists as an average of the two.

**Polar Molecule** - A molecule that is mostly positive on one side and mostly negative on the other.

**$sp^3$  Orbital** - An orbital composed of one s orbital and three p orbitals; can hold eight total electrons. Tetrahedral.

**Trigonal Planar** - The way that three objects will arrange themselves around a central atom. Non-polar.

**Free Radicals** - Compounds with an odd number of electrons in their Lewis Structure. Eg.  $\text{NO}_2$

**Incomplete Octets** - Occur in elements that can form stable bonds with less than eight electrons. Eg.  $\text{BF}_3$

**Expanded Octets** - Occur in the third row of the periodic table and beyond in elements that can form stable bonds with more than eight electrons. Eg.  $\text{XeF}_2$

**Hydrogen Bonding** - When hydrogen is attracted to electronegative atoms.

# Acs General Chemistry Exam Study Guide

**IM Harris**



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## **Table of Contents Acs General Chemistry Exam Study Guide**

1. Understanding the eBook Acs General Chemistry Exam Study Guide
  - The Rise of Digital Reading Acs General Chemistry Exam Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Acs General Chemistry Exam Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acs General Chemistry Exam Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acs General Chemistry Exam Study Guide
  - Personalized Recommendations
  - Acs General Chemistry Exam Study Guide User Reviews and Ratings
  - Acs General Chemistry Exam Study Guide and Bestseller Lists
5. Accessing Acs General Chemistry Exam Study Guide Free and Paid eBooks
  - Acs General Chemistry Exam Study Guide Public Domain eBooks
  - Acs General Chemistry Exam Study Guide eBook Subscription Services
  - Acs General Chemistry Exam Study Guide Budget-Friendly Options

6. Navigating Acs General Chemistry Exam Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Acs General Chemistry Exam Study Guide Compatibility with Devices
  - Acs General Chemistry Exam Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acs General Chemistry Exam Study Guide
  - Highlighting and Note-Taking Acs General Chemistry Exam Study Guide
  - Interactive Elements Acs General Chemistry Exam Study Guide
8. Staying Engaged with Acs General Chemistry Exam Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Acs General Chemistry Exam Study Guide
9. Balancing eBooks and Physical Books Acs General Chemistry Exam Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Acs General Chemistry Exam Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Acs General Chemistry Exam Study Guide
  - Setting Reading Goals Acs General Chemistry Exam Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Acs General Chemistry Exam Study Guide
  - Fact-Checking eBook Content of Acs General Chemistry Exam Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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