



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{KNO}_3$
  - $\text{H}_2\text{SO}_4$
  - $\text{H}_2\text{O}_2$
  - $\text{C}_2\text{H}_6$
  - $\text{H}_2$
  - $\text{O}_2$

- Classify each of the following as either ionic or molecular compound. Write the name of the compound.
- $\text{K}_2\text{SO}_4$
- $\text{C}_2\text{H}_6$
- $\text{H}_2\text{O}$
- $\text{H}_2\text{SO}_4$
- $\text{H}_2\text{O}_2$
- $\text{C}_2\text{H}_6$
- $\text{H}_2$
- $\text{O}_2$

### SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride.  $\text{H}-\text{F}$
- Show covalent bonding in methane.  $\text{C}$  with four  $\text{H}$  atoms around it.
- Show covalent bonding in carbon dioxide.  $\text{C}$  with two  $\text{O}$  atoms around it.
- Show covalent bonding in ethane.  $\text{H}-\text{C}-\text{C}-\text{H}$  with appropriate  $\text{H}$  atoms.
- Show covalent bonding in acetylene.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for the following molecules.
  - $\text{H}_2\text{O}$
  - $\text{H}_2\text{O}_2$
  - $\text{C}_2\text{H}_6$
  - $\text{C}_2\text{H}_4$
  - $\text{C}_2\text{H}_2$
  - $\text{CO}_2$
  - $\text{SO}_2$
  - $\text{NO}_2$
  - $\text{NH}_3$
  - $\text{H}_2\text{S}$
  - $\text{H}_2\text{Se}$
  - $\text{H}_2\text{Te}$
  - $\text{H}_2\text{Po}$
  - $\text{H}_2\text{O}$
  - $\text{H}_2\text{O}_2$
  - $\text{C}_2\text{H}_6$
  - $\text{C}_2\text{H}_4$
  - $\text{C}_2\text{H}_2$
  - $\text{CO}_2$
  - $\text{SO}_2$
  - $\text{NO}_2$
  - $\text{NH}_3$
  - $\text{H}_2\text{S}$
  - $\text{H}_2\text{Se}$
  - $\text{H}_2\text{Te}$
  - $\text{H}_2\text{Po}$

### SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
  - $5s$
  - $5p$
  - $6s$
  - $6p$
  - $7s$
  - $7p$
- Which orbital and sub-orbital are represented by each of the following?
  - $1s$
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
  - $5s$
  - $5p$
  - $6s$
  - $6p$
  - $7s$
  - $7p$

# Chapter 8 Covalent Bonding Practice Problems Answers Bing

**M Mark**



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