



# COVALENT BONDING

## Practice Problems

Do your answers to the following problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either ionic or molecular.
  - $\text{NaCl}$
  - $\text{H}_2\text{O}$
  - $\text{HCl}$
  - $\text{KNO}_3$
  - $\text{H}_2\text{SO}_4$
  - $\text{H}_2\text{O}_2$
- Write the formula for the following molecular compounds:
  - nitrogen dioxide
  - carbon dioxide
  - carbon monoxide
  - nitrogen monoxide
  - nitrogen trioxide
  - nitrogen pentoxide

- Write the molecular formula for each of the following compounds:
  - nitrogen dioxide
  - carbon dioxide
  - carbon monoxide
  - nitrogen monoxide
  - nitrogen trioxide
  - nitrogen pentoxide
- Write the molecular formula for each of the following compounds:
  - nitrogen dioxide
  - carbon dioxide
  - carbon monoxide
  - nitrogen monoxide
  - nitrogen trioxide
  - nitrogen pentoxide

### SECTION 8.2 THE NATURE OF COVALENT BONDING

- Show covalent bonding in hydrogen fluoride.  $\text{H}-\text{F}$
- Show covalent bonding in methane.  $\text{C}$  with four  $\text{H}$  atoms around it.
- Show covalent bonding in carbon dioxide.  $\text{O}=\text{C}=\text{O}$
- Show covalent bonding in acetylene.  $\text{H}-\text{C}\equiv\text{C}-\text{H}$
- Draw Lewis structures for the following:
  - methane
  - ammonia
  - water
  - hydrogen peroxide
  - carbon dioxide
  - nitrogen dioxide
  - nitrogen monoxide
  - nitrogen trioxide
  - nitrogen pentoxide



### SECTION 8.3 BONDING: THE ORBITALS

- Which orbital and sub-orbital are represented by each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
  - $5s$
  - $5p$
- Which orbital and sub-orbital are represented by each of the following?
  - $2s$
  - $2p$
  - $3s$
  - $3p$
  - $4s$
  - $4p$
  - $5s$
  - $5p$

# Chapter 8 Covalent Bonding Test B Answers

**N Noddings**



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