



COVALENT BONDING

Practice Problems

Do your best to solve each of the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

- Classify each of the following as either an ionic or a molecular compound.
 - NaCl
 - H_2O
 - CaF_2
 - C_2H_6
 - KNO_3
 - H_2S
 - MgO
 - C_2H_4
 - H_2SO_4
 - H_2O_2
 - C_2H_2
 - $\text{C}_2\text{H}_6\text{O}$
 - $\text{C}_2\text{H}_4\text{O}_2$
 - $\text{C}_2\text{H}_2\text{O}_2$
 - $\text{C}_2\text{H}_2\text{Cl}_2$
 - $\text{C}_2\text{H}_2\text{Cl}_4$
 - $\text{C}_2\text{H}_2\text{Cl}_6$
- Write the formula for the following molecular compounds:
 - nitrogen dioxide
 - nitrogen monoxide
 - carbon dioxide
 - carbon monoxide
 - diphosphorus pentoxide
 - diphosphorus trioxide
 - diphosphorus tetroxide
 - diphosphorus hexoxide
 - diphosphorus heptoxide
 - diphosphorus octoxide
 - diphosphorus nonoxide
 - diphosphorus decaoxide
 - phosphorus pentachloride
 - phosphorus pentabromide
 - phosphorus pentaiodide
 - phosphorus pentasulfide
 - phosphorus pentaselenide
 - phosphorus pentatelluride
 - phosphorus pentastibide
 - phosphorus pentabismuthide
 - phosphorus pentatermide
 - phosphorus pentastannide
 - phosphorus pentaplumbide
 - phosphorus pentabismuthide
 - phosphorus pentatermide
 - phosphorus pentastannide
 - phosphorus pentaplumbide
- Write the name for each of the following molecular compounds:
 - N_2O
 - NO
 - NO_2
 - N_2O_3
 - N_2O_4
 - N_2O_5
 - P_2O_3
 - P_2O_4
 - P_2O_5
 - P_2O_6
 - P_2O_7
 - P_2O_8
 - P_2O_9
 - P_2O_{10}
 - PCl_5
 - PBr_5
 - PI_5
 - P_2S_5
 - P_2Se_5
 - P_2Te_5
 - P_2Sb_5
 - P_2Bi_5
 - P_2Tl_5
 - P_2Sn_5
 - P_2Pb_5
 - P_2Bi_5
 - P_2Tl_5
 - P_2Sn_5
 - P_2Pb_5

SECTION 8.2 THE NATURE OF COVALENT BONDING

- Draw the Lewis structure for hydrogen fluoride. $\text{H}-\text{F}$
- Draw the Lewis structure for hydrogen chloride. $\text{H}-\text{Cl}$
- Draw the Lewis structure for hydrogen bromine. $\text{H}-\text{Br}$
- Draw the Lewis structure for hydrogen iodine. $\text{H}-\text{I}$
- Draw the Lewis structure for hydrogen cyanide. $\text{H}-\text{C}\equiv\text{N}$
- Draw the Lewis structure for hydrogen cyanide. $\text{H}-\text{C}\equiv\text{N}$



SECTION 8.3 EXERCISES: THE CARBON

- Write the Lewis structure for each of the following:
 - methane
 - ethane
 - propane
 - butane
 - pentane
 - hexane
 - heptane
 - octane
 - nonane
 - decane
- Write the Lewis structure for each of the following:
 - ethene
 - propene
 - butene
 - pentene
 - hexene
 - heptene
 - octene
 - nonene
 - decene

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Padhraic Smyth



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