

Chapter 9 Practice Problems: Chemical Names and Formulas

Section 9.1: Naming Ions

- What is the charge on the ion typically formed by each element?
 - oxygen
 - iodine
 - sodium
 - aluminum
 - nickel, two electrons lost
 - magnesium
- How many electrons does the neutral atom gain or lose when each ion forms?
 - Cr^{3+}
 - P^{3-}
 - Li^+
 - Ca^{2+}
 - Cl^-
 - O^{2-}
- Name each ion. Identify each as a cation or an anion.
 - Sn^{2+}
 - Co^{3+}
 - Br^-
 - K^+
 - H^-
 - Mn^{2+}
- Write the formula (including the charge) for each ion. Use table 9.3 as if necessary.
 - carbonate ion
 - nitrite ion
 - sulfate ion
 - hydroxide ion
 - chromate ion
 - ammonium ion
- Name each of the following ions. Identify each as a cation or an anion.
 - CN^-
 - HCO_3^-
 - PO_4^{3-}
 - Cl^-
 - Ca^{2+}
 - SO_3^{2-}

9.2 Naming and Writing Formulas for Ionic Compounds:

- Write the formulas for these binary ionic compounds.
 - magnesium oxide
 - tin(II) fluoride
 - potassium iodide
 - aluminum chloride
 - sodium sulfide
 - ferric bromide
- Write the formulas for the compounds formed from these pairs of ions.
 - Ba^{2+} and Cl^-
 - Ag^+ and I^-
 - Ca^{2+} and S^{2-}
 - K^+ and Br^-
 - Al^{3+} and O^{2-}
 - Fe^{2+} and O^{2-}
- Name the following binary ionic compounds.
 - MnO_2
 - Li_3N
 - CaCl_2
 - SrBr_2
 - NiCl_2
 - K_2S
 - Fe_3N_2
 - SnCl_4
- Write formulas for the following ionic compounds.
 - sodium phosphate
 - magnesium sulfate
 - sodium hydroxide
 - potassium cyanide
 - ammonium chloride
 - potassium dichromate
- Write formulas for the following ionic compounds.
 - NH_4^+ and SO_4^{2-}
 - K^+ and NO_3^-
 - barium ion and hydroxide ion
 - lithium ion and carbonate ion

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Understanding Chemistry Fred M. Dewey, 1994 **General Organic and Biological Chemistry** Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed Peter P. Berlow, 1978 *Addison-Wesley Chemistry*, 1990 *Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Scientific Soapmaking*

Kevin M. Dunn,2010 Scientific Soapmaking bridges the gap between the technical and craft literature It explains the chemistry of fats oils and soaps and teaches sophisticated analytical techniques that can be carried out using equipment and materials familiar to makers of handcrafted soap *The Practice of Chemistry* Donald J. Wink,Sharon Fetzer-Gislason,Sheila McNicholas,2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POC/PowerPoint_Final.ppt 832KB **Partial Solutions Manual** Darrell D. Ebbing,David Bookin,1998 *Study Guide/Selected Solutions Manual* Julie R. Frentrup,2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book *Prentice Hall Chemistry*,2000 *New Technical Books* New York Public Library,1925 *Chemical Engineering Catalog*,1921 *The Digester, Over the Spillway*,1976 *Basic Concepts of Chemistry* Leo J. Malone,Theodore Dolter,2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance **Chemistry** James E. Brady,Fred Senese,2004-02-04 Publisher Description **Farmer's Advocate and Home Magazine**,1897 **Scientific American**,1888 [Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science](#),2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction

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Table of Contents Chapter 9 Chemical Names And Formulas Practice Problems Answers

1. Understanding the eBook Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - The Rise of Digital Reading Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - Personalized Recommendations
 - Chapter 9 Chemical Names And Formulas Practice Problems Answers User Reviews and Ratings
 - Chapter 9 Chemical Names And Formulas Practice Problems Answers and Bestseller Lists
5. Accessing Chapter 9 Chemical Names And Formulas Practice Problems Answers Free and Paid eBooks
 - Chapter 9 Chemical Names And Formulas Practice Problems Answers Public Domain eBooks
 - Chapter 9 Chemical Names And Formulas Practice Problems Answers eBook Subscription Services

- Chapter 9 Chemical Names And Formulas Practice Problems Answers Budget-Friendly Options
- 6. Navigating Chapter 9 Chemical Names And Formulas Practice Problems Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 9 Chemical Names And Formulas Practice Problems Answers Compatibility with Devices
 - Chapter 9 Chemical Names And Formulas Practice Problems Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - Highlighting and Note-Taking Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - Interactive Elements Chapter 9 Chemical Names And Formulas Practice Problems Answers
- 8. Staying Engaged with Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 9 Chemical Names And Formulas Practice Problems Answers
- 9. Balancing eBooks and Physical Books Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 9 Chemical Names And Formulas Practice Problems Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - Setting Reading Goals Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - Fact-Checking eBook Content of Chapter 9 Chemical Names And Formulas Practice Problems Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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