

Section 12.1

OBJECTIVE 1

Write complete balanced molecular equations that are indicated.

1. Write the balanced molecular equation for each reaction that occurs.
 - (a) H_2SO_4 and NaOH in aqueous solution
 - (b) H_2SO_4 and Ca(OH)_2 in aqueous solution
 - (c) H_2SO_4 and CaCO_3 in aqueous solution
 - (d) H_2SO_4 and CaCO_3 in aqueous solution
2. Write the balanced molecular equation for each reaction that occurs.
 - (a) H_2SO_4 and CaCO_3 in aqueous solution
 - (b) H_2SO_4 and CaCO_3 in aqueous solution
 - (c) H_2SO_4 and CaCO_3 in aqueous solution
 - (d) H_2SO_4 and CaCO_3 in aqueous solution
3. Write the balanced molecular equation for each reaction that occurs.
 - (a) H_2SO_4 and CaCO_3 in aqueous solution
 - (b) H_2SO_4 and CaCO_3 in aqueous solution
 - (c) H_2SO_4 and CaCO_3 in aqueous solution
 - (d) H_2SO_4 and CaCO_3 in aqueous solution
4. Write the balanced molecular equation for each reaction that occurs.
 - (a) H_2SO_4 and CaCO_3 in aqueous solution
 - (b) H_2SO_4 and CaCO_3 in aqueous solution
 - (c) H_2SO_4 and CaCO_3 in aqueous solution
 - (d) H_2SO_4 and CaCO_3 in aqueous solution

5. Write the balanced molecular equation for each reaction that occurs.



(a) Write the equation:



(b) Write the balanced molecular equation for the reaction.

(c) Write the balanced molecular equation for the reaction.

(d) Write the balanced molecular equation for the reaction.

(e) Write the balanced molecular equation for the reaction.

Chapter 9 Section 1 Review Stoichiometry Answers

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**

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