

## Laboratory Report – Periodic Table

Type Name Here:

### Part A

1. Place 9 clean small test tubes in a test tube rack. The test tubes should be clean but do not have to be dry because aqueous solutions will be used.
2. To each test tube add 5 drops of the solution (0.1 M) indicated in the data table.
3. Next add 5 drops of silver nitrate (0.5 M) to each of the 9 test tubes. If there is no apparent change be sure to gently swirl the test tube to mix the reagents (a stirring rod can be used). If the contents of the test tube are cloudy, a solid precipitate was formed. If the contents of the test tube are clear, no reaction occurred. Record the results in the data table writing "ppt" and their color for precipitates and "NR" for no reaction.
4. Your instructor will inform you of where to dispose of the chemicals used. Rinse the test tubes and repeat the procedure two more times first using ammonium carbonate (0.5 M) in step 3 and then again using sodium phosphate (0.5 M) in step 3.

Test Tube	Solution	Test Ion	Ag <sup>+</sup>	CO <sub>3</sub> <sup>2-</sup>	PO <sub>4</sub> <sup>3-</sup>
1	Barium nitrate	Ba <sup>2+</sup>		White ppt	White ppt
2	Ammonium chloride	Cl <sup>-</sup>	White ppt		
3	Calcium nitrate	Ca <sup>2+</sup>		White ppt	White ppt
4	Lithium nitrate	Li <sup>+</sup>			
5	Ammonium iodide	I <sup>-</sup>	Yellowish ppt		
6	Potassium nitrate	K <sup>+</sup>			
7	Sodium nitrate	Na <sup>+</sup>			
8	Strontium nitrate	Sr <sup>2+</sup>		White ppt	White ppt
9	Ammonium bromide	Br <sup>-</sup>	Off-White ppt		

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