

**WORKSHEET**  
**CHEMISTRY FOR ENGINEERING STUDENTS, LECTURE 1**  
 Molecular formula of hydrocarbons and the structural formula of hydrocarbons

**ANSWERS KEY**

**QUESTION 1. Calculate the empirical formula of hydrocarbons and alcohols.**

1. Determination of formula of hydrocarbons based on percentage composition of carbon and hydrogen. **Example:**  $\text{C}_2\text{H}_6$  or  $\text{C}_3\text{H}_8$ . These hydrocarbons are formed by combustion of the hydrocarbon.
  - (a) write the balanced chemical equation for the reaction:  
 $2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$
  - (b) compare the mass of the hydrocarbon with the mass of carbon and hydrogen in the products.
  - (c) find the mass of the hydrocarbon from the mass of carbon and hydrogen.
  
2. Same process as above but with oxygen included to determine the molecular formula of alcohols. **Example:**  $\text{C}_2\text{H}_6\text{O}$  or  $\text{C}_3\text{H}_8\text{O}$ . These hydrocarbons are formed by the combustion of the hydrocarbon.
  - (a) write the balanced chemical equation for the reaction:  
 $2\text{C}_2\text{H}_6\text{O} + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$
  - (b) compare the mass of the hydrocarbon with the mass of carbon, hydrogen and oxygen in the products.
  - (c) find the mass of the hydrocarbon from the mass of carbon, hydrogen and oxygen.
  
3. Determination of formula of alcohols based on percentage composition of carbon, hydrogen and oxygen. **Example:**  $\text{C}_2\text{H}_6\text{O}$  or  $\text{C}_3\text{H}_8\text{O}$ .
  - (a) write the balanced chemical equation for the reaction:  
 $2\text{C}_2\text{H}_6\text{O} + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$
  - (b) find the mass of the hydrocarbon from the mass of carbon, hydrogen and oxygen.
  - (c) find the mass of the hydrocarbon from the mass of carbon, hydrogen and oxygen.

**QUESTION 2. Calculate the structural formula of hydrocarbons and alcohols.**

1. Draw the structural formula of hydrocarbons and alcohols based on the molecular formula. **Example:**  $\text{C}_2\text{H}_6$  or  $\text{C}_3\text{H}_8$ .
  - (a) write the balanced chemical equation for the reaction:  
 $2\text{C}_2\text{H}_6 + 7\text{O}_2 \rightarrow 4\text{CO}_2 + 6\text{H}_2\text{O}$
  - (b) compare the mass of the hydrocarbon with the mass of carbon and hydrogen in the products.

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