

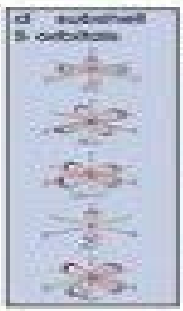
CHEMISTRY

OPEN

PERIODIC TABLE OF THE ELEMENTS

Diagram showing the periodic table of elements, color-coded by groups. Hydrogen (H) is highlighted with a callout box:

- 1** - Atomic Number
- H** - Atomic Symbol
- 1.008** - Atomic Weight



ATOMIC STRUCTURE

Atomic Number, Z , of an element is the number of protons in the nucleus. For a neutral atom, $Z = +$ the number of electrons. The mass number, A , is the sum of the number of protons and neutrons. Isotopes of an element have the same Z , different A . The atomic mass is the weighted average of isotopes. See also: Atomic Structure, Atomic Weight, Crystal Structure of Solids, Electron Spin, Ionization Energy, Periodic Table of Elements.

ATOMIC QUANTUM NUMBERS AND ORBITALS

Principal quantum number, n , is the number of shells. $n = 1, 2, 3, \dots$. Subshells are labeled s, p, d, f . Each shell has a certain number of orbitals. For example, $n=2$ has 4 orbitals (1 s , 3 p). Each orbital can hold 2 electrons.

NUCLEAR CHEMISTRY

Nuclear reactions occur when particles (neutrons, protons, alpha particles, beta particles) strike a nucleus, causing it to change. Energy is released or absorbed. Radioactive decay involves the transformation of an unstable nucleus into a stable one. Types of decay: α , β , γ , positron emission, electron capture. Half-life is the time for half of a substance to decay.

MANY-ELECTRON ATOMS

Orbitals and quantum numbers are used to describe the arrangement of electrons in an atom. The Aufbau principle states that electrons fill orbitals from lowest to highest energy. Hund's rule states that electrons fill degenerate orbitals singly before pairing up. Pauli exclusion principle states that no two electrons can have the same four quantum numbers.

n	l	m _l	m _s	Orbitals
1	0	0	$\pm \frac{1}{2}$	1 orbital
2	0	0	$\pm \frac{1}{2}$	1 orbital
2	1	1, 0, -1	$\pm \frac{1}{2}$	3 orbitals
3	0	0	$\pm \frac{1}{2}$	1 orbital
3	1	1, 0, -1	$\pm \frac{1}{2}$	3 orbitals
3	2	2, 1, 0, -1, -2	$\pm \frac{1}{2}$	5 orbitals
4	0	0	$\pm \frac{1}{2}$	1 orbital
4	1	1, 0, -1	$\pm \frac{1}{2}$	3 orbitals
4	2	2, 1, 0, -1, -2	$\pm \frac{1}{2}$	5 orbitals
4	3	3, 2, 1, 0, -1, -2, -3	$\pm \frac{1}{2}$	7 orbitals

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