

QUALITATIVE ANALYSIS

Experiment	Observation	Inference
<u>Preliminary Test</u>		
1. Noted the colour of the given salt	<ol style="list-style-type: none"> Blue or Bluish Green Greenish Light Green Dark Brown Light Pink Colourless 	<p>Cu^{2+} or Ni^{2+} Ni^{2+} Fe^{2+} Fe^{3+} Mn^{2+} Absence of Cu^{+2}, Ni^{2+}, Fe^{2+}, Fe^{3+}, Mn^{2+}</p>
2. Noted the solubility	<ol style="list-style-type: none"> Soluble in water Soluble in dilute HCl 	<p>NH_4^+ may be present Absence of group I cations</p>
3. Dry Heating Test: Heated a pinch of the salt in a dry Test tube	<ol style="list-style-type: none"> Colour, odourless gas Colourless gas with rotten egg smell Colourless gas with vinegar smell Colourless gas with ammonia smell Reddish brown gas Greenish yellow gas Violet gas 	<p>CO_3^{2-} may be present S^{2-} may be present CH_3COO^- may be present NH_4^+ may be present NO_3^- or Br^- may be present Cl^- may be present I^- may be present</p>
<u>Identification of acid radical (Anion)</u>		
1. Dilute Acid Test: Treated a pinch of the substance With dilute H_2SO_4	<ol style="list-style-type: none"> Colourless, odourless gas with brisk effervescence which turned lime water milky. Vinegar Smell No characteristic Reaction 	<p>CO_3^{2-} present $\text{CO}_3^{2-} + 2\text{H}^+ \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ $\text{Ca}(\text{OH})_2 + \text{CO}_2 \rightarrow \text{CaCO}_3 + \text{H}_2\text{O}$</p> <p>$\text{CH}_3\text{COO}^-$ Present $\text{CH}_3\text{COO}^- + \text{H}^+ \rightarrow \text{CH}_3\text{COOH}$</p> <p>Absence of CO_3^{2-}, CH_3COO^-</p>
2. Concentrated Acid Test: Treated a pinch of the substance With concentrated H_2SO_4	<ol style="list-style-type: none"> Colourless gas with pungent Smell which gave dense white fumes when a glass rod dipped in NH_4OH was brought near the mouth of the test tube Reddish brown vapours with Pungent smell. Solution turns red Deep violet vapours with pungent smell. Sides of test tube 	<p>Cl^- present $\text{Cl}^- + \text{H}^+ \rightarrow \text{HCl}$ $\text{HCl} + \text{NH}_4\text{OH} \rightarrow \text{NH}_4\text{Cl} + \text{H}_2\text{O}$</p> <p>$\text{Br}^-$ present $\text{Br}^- + \text{H}^+ \rightarrow \text{HBr}$ $2\text{HBr} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Br}_2 + \text{SO}_2$</p> <p>$\text{I}^-$ present $\text{I}^- + \text{H}^+ \rightarrow \text{HI}$ $2\text{HI} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{I}_2 + \text{SO}_2$</p>

Chemistry Practical Qualitative Analysis Sheet

Zhe Xu



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