

Nomenclature Review

Name: \_\_\_\_\_

1. Give the charge and name for each of the following *oxyanions*.

	Charge	Name
a. $\text{NO}_3^-$	-1	nitrate
b. $\text{CO}_3^{2-}$	-2	carbonate
c. $\text{C}_2\text{H}_3\text{O}_2^-$	-1	acetate
d. $\text{ClO}_2^-$	-1	chlorite
e. $\text{PO}_3^{2-}$	-3	phosphite
f. $\text{ClO}_4^-$	-1	perchlorate
g. $\text{SO}_4^{2-}$	-2	sulfate
h. $\text{NH}_4^+$	+1	ammonium
i. $\text{OH}^-$	-1	hydroxide
j. $\text{CN}^-$	-1	cyanoide
k. $\text{BrO}_3^-$	-1	bromate
l. $\text{C}_2\text{O}_4^{2-}$	-2	oxalate
m. $\text{SO}_3^{2-}$	-2	sulfite
n. $\text{MnO}_4^-$	-1	permanganate

2. If A and B represent compounds of an element bonded to a polyatomic ion, give an example of a valid formula for each combination below.

a. AB	b. $\text{AB}_2$	c. $\text{A}_2\text{B}$	d. $\text{AB}_3$	e. $\text{A}_2\text{B}_3$
$\text{Al}(\text{NO}_3)_3$	$\text{Ca}(\text{NO}_2)_2$	$(\text{NH}_4)_2\text{S}$	$\text{Al}(\text{ClO}_4)_3$	$\text{Al}_2(\text{SO}_4)_3$

3. What is the oxidation number of:

a. Calcium ion?	+2		
b. Sodium ion?	+1		
c. Sulfur ion?	-2		
d. Phosphorus ion?	-3		
e. Carbon ion within Carbon dioxide molecule?	+4	$x + 2(-2) = 0$ $x = 4+$	$\text{CO}_2$
f. Phosphorus within the Phosphate polyatomic ion?	+5		$\text{PO}_4^{3-}$ $x + 4(-2) = -3$ $x = +6$
g. Chromium within the Chromate polyatomic ion?	+6		
h. Carbon within the Oxalate polyatomic ion?	+3		$\text{C}_2\text{O}_4^{2-}$ $2x + 4(-2) = -2$ $x = 3+$
i. Zinc ion?	+2		
j. Cadmium ion?	+2		
k. Sulfur within the Hydrogen Sulfite ion ( $\text{HSO}_3^-$ )	+4	$1 + x - 6 = -1$ $x = 4+$	

4. Give the most common oxidation number/numbers for each of the following groups or families of elements.

a. Alkali Metals	+1
b. Alkaline Earth	+2
c. Noble Gases	0
d. Halogens	-1
e. Group 13	+3
f. Group 15	-3

5. Give the formula for the following hydrocarbons/alcohols:

a. methane	$\text{C}_1\text{H}_4$
b. butene	$\text{C}_4\text{H}_8$
c. ethanol	$\text{C}_2\text{H}_5\text{OH}$
d. propyne	$\text{C}_3\text{H}_4$

# Chemistry Unit 4 Test Review Electron Configuration

**Michael Rajnik**



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